Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
 - a. Write the balanced equation for the reaction given above:

+ NaNO₁ → Cu(NO₁)₂ + ChrCile CuCl₂ + 2 NaNO₃ → Cu(NO₃)₂ + 2 NaCl

 If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

1 mol CuCl₂ x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 15 g CuClax 134.362g CuCl: 1 mol CuCl: 1 mol NaCl

20 g NaNO₃ x 1 mol NaNO₃ x 2 mol NaCl x 58.443 g NaCl = 13.8 g NaCl 84,997g NaNOs 2 mol NaNOs 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride
- d. How many grams of copper (II) nitrate is formed?

 $15g CuCl_2 \times 1 \mod CuCl_2 \times 1 \mod Cu(NO_3)_2 \times 187.467 g Cu(NO_3)_2 = 20.9 g Cu(NO_3)_2$ 134,452g CuCl₂ 1 mol CuCl₂ 1 mol Cu(NO₃)₂

e. How much of the excess reasent is left over in this reaction?

15g CuCl₂ x 1 mol CuCl₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 19 g NaNO₃ 134.452g CuCl₂ 1 mol CuCl₂ 1 mol NaNO₃

20 g NaNOs - 19 g NaNOs = 1 g NaNOs

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is Percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$ the percent yield of this reaction?

formed.

11.3/13.0 x 100% = 86.9%

- 2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are
 - a. Balance the following equation:

Pb(NO₃)_{2 (and} + NaI (and → PbI_{2 (a)} + NaNO_{3 (and} Pb(NO₃)_{2 tant} + 2 NaI _{tant} → PbI_{2 tot} + 2 NaNO_{3 tant}

 If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO₃)₂ x 1 mol Pb(NO₃)₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 12.8 g NaNO₃ 331.214 g Pb(NO₃)₂ 1 mol Pb(NO₃)₂ 1 mol NaNO₃

15.0 g Nal x 1 mol Nal x 2 mol NaNO₃ x 84.997g NaNO₃ = 8.51 g NaNO₃ 149,89g Nat 2 mol NaI 1 mol NaNOs

8.51 g NaNO3 can be made

<u>Limiting Reagent And Percentage Yield Homework</u> <u>Answers</u>

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