

## CHEMICAL BONDING VOCABULARY:

### **Chemical Bonding:**

a mutual electrical attraction between the **nuclei** (*positive*) and **valence electrons** (*negative*) of **different** atoms that binds the atoms together.

- interaction that holds 2 atoms together.
- This forms a **NEW substances** with **NEW properties**.

**Why bond?** Most atoms are not stable by themselves.

A chemical bond provides **stability** for both atoms involved by lowering the potential energy of both atoms.

**Stability:** Atoms want **8 valence electrons** for maximum stability (like a noble gas), so they either lose, gain, or share electrons from their outermost **energy levels** to get 8.

**Exception:** Helium has only 2 valence electrons in its 2nd Level which makes its outer level full & stable.

**Compounds:** atoms of different elements bonded together

**Binary Compounds:** Chemical compounds made of two distinct elements.

#### Two Types:

- **Ionic Compounds:** Compounds made of ions.
- **Covalent Compounds:** Chemical bond in which two atoms share valence electrons.

**Charge Balance** -The sum of positive charges equals the sum of negative charges in solution (Electroneutrality). **ONLY** for **IONIC** Compounds.

**Subscripts** are numbers that come after a symbol and below. Subscripts tell you the number of atoms of that element. If an element does not have a subscript, then it is understood that the subscript is 1.

**A chemical nomenclature** is a set of rules to generate systematic names for chemical compounds. The nomenclature used most frequently worldwide is the one created and developed by the International Union of Pure and Applied Chemistry (IUPAC).

# L1 Chemical Bonding Vocab Activity

**R Sanford**



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