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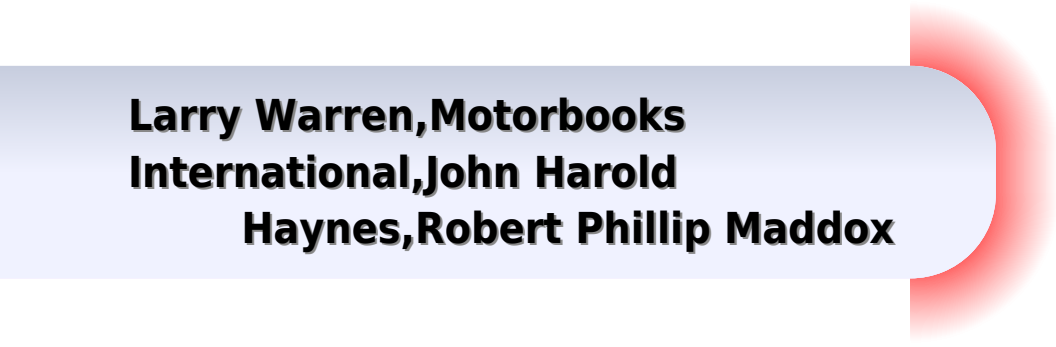
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Isuzu Rodeo Workshop Manual

Repair, Service, Maintenance, Wiring Diagrams etc

Isuzu Rodeo 2002 Factory Service Repair Manual

**Larry Warren, Motorbooks
International, John Harold
Haynes, Robert Phillip Maddox**



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Automotive News, 2000

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1998 Isuzu Rodeo (UE) Workshop Manual Isuzu Jidōsha Kabushiki Kaisha, 1997

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Isuzu Trooper & Pick-up Automotive Repair Manual Larry Warren, Motorbooks International, John Harold Haynes, Robert Phillip Maddox, 1990-07-01

Workshop Manual Isuzu Jidōsha Kabushiki Kaisha, 1994

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Workshop Manual, MY00 Rodeo Service Manual Isuzu Jidōsha Kabushiki Kaisha, 2000*

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Isuzu Workshop Manual 1993 Supplement, 1993

Isuzu Truck Repair Manual, 1980

Rodeo (UCR/UCS) 1997 Workshop Manual, 1996

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 $\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$
 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $\text{Pb} + \text{OH}^- \rightarrow \text{PbCl}_2 + \text{H}_2\text{O}$
 $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$
 $2\text{KBr} + \text{Al}_2\text{SO}_4 \rightarrow 2\text{K}_2\text{SO}_4 + 3\text{Br}_2$
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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $2\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{NaBr} \rightarrow \text{Na}_3\text{GaO}_3 + \text{SO}_3$
 $3\text{NF}_3 + 5\text{SnF}_2 \rightarrow \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr}$
 $2\text{NF}_3 + 3\text{SnF}_2 \rightarrow \text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$
how many moles of Ce ZnCl_2 Zn will be produced from 23.0 g of Ce 23.0 g of Zn

assuming CeCl_2 CuCl_2 is available in excess

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 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{Na}_2\text{SO}_4$
br using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{O}_3 + \text{SO}_3 + \text{NaBr}$
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same thing goes with the o3 i am so lost

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