



# La Veritafa Sui Restauri Bolognesi

**Guido Zucchini**



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measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less  
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of  
sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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eg 1 HCl aq + NaOH aq → NaCl aq + H<sub>2</sub>O l for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub> aq + KOH aq → KNO<sub>3</sub> aq + H<sub>2</sub>O l for which ΔH = 57 kJ mol<sup>-1</sup>

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together to produce one mole of water for example HCl aq + NaOH aq → NaCl aq + H<sub>2</sub>O aq you can see that one mole of water is  
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