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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{\ln \left(\frac{m_0}{m} \right)}{\ln 2}$ $t = \frac{n \times t_{1/2}}{\ln 2}$ $x = 50.3$
 125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{\ln \left(\frac{100}{80} \right)}{\ln 2}$ $t = 80$ $n = 100$
 now using the half life equation in example 1 we have

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