## 16 • Chemical Equilibrium

Consider the reaction system,

 $CoO(s) + H_2(g) \stackrel{\text{def}}{\leftarrow} Co(s) + H_2O(g)$ .

The equilibrium constant expression is

- [CoO][H.] [Co][H.O]
- [H.O]
- $[Co][H_2O]$ [CoO][H.] Col[H.O] HI.
- Given the equilibrium.

2SO<sub>2</sub>(g) + O<sub>2</sub>(g) ∓ 2SO<sub>2</sub>(g), if this equilibrium

is established by beginning with equal number of moles of SO<sub>2</sub> and O<sub>2</sub> in a 1.0 Liter bulb, then

the following must be true at equilibrium:

- a) [SO<sub>2</sub>] = [SO<sub>3</sub>]d) [SO<sub>2</sub>] < [O<sub>2</sub>]
- (b)  $2|SO_2| = 2|SO_3|$  e)  $|SO_2| > |O_2|$ 
  - c) [SO<sub>2</sub>] = [O<sub>2</sub>]

#### Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO, 0.200 mol Cb and 0.500 mol CINO were placed in a 25.0 Liter container. The following equilibrium is established: 2CINO(g) ₹ 2NO(g) + Cb(g)

- At equilibrium, 0.600 mol of CINO was present. The number of moles of Cb present at equilibrium is
  - cal 0.050

e) 0.250

b) 0.100 c) 0.150

- PRACTICE TEST
- The equilibrium constant, Kc, is:
  - a) 4.45 x 10<sup>-4</sup>
- d) 0.167
- b) 6.67 x 10<sup>-4</sup>
- e) 1500

- ch 0.111
- At 985°C, the equilibrium constant for the reaction.

$$H_2(g) + CO_2(g) \rightleftharpoons H_2O(g) + CO(g)$$

is 1.63. What is the equilibrium constant for the reverse reaction?

a) 1.63

d) 0.613

b) 0.815

e)) 1.00

- c) 2.66
- What is the relationship between K<sub>p</sub> and K<sub>c</sub> for the reaction,  $2ICI(g) \rightleftharpoons I_2(g) + Cb(g)$ ?
  - (a)  $k_p = K_c(RT)^{-1}$  (d)  $K_p = K_c$

- b)  $K_p = K_c(RT)$  e)  $K_p = K_c(2RT)$
- c)  $K_p = K_c(RT)^2$
- For the reaction 2NO<sub>2</sub>(g) ₹ N<sub>2</sub>O<sub>4</sub>(g), K<sub>n</sub> at 25°C is 7.3, when all partial pressures are expressed in atmospheres. What is Ke for this reaction? [R=0.0821 Latmmol\* K\*]

4270

d) 179

b) 0.0119

e) 2.06

c) = 0.291

# **Mixed Review Chemical Equilibrium Answers**

**K Payea** 

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## **Table of Contents Mixed Review Chemical Equilibrium Answers**

- 1. Understanding the eBook Mixed Review Chemical Equilibrium Answers
  - The Rise of Digital Reading Mixed Review Chemical Equilibrium Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Mixed Review Chemical Equilibrium Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mixed Review Chemical Equilibrium Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mixed Review Chemical Equilibrium Answers
  - Personalized Recommendations
  - Mixed Review Chemical Equilibrium Answers User Reviews and Ratings
  - Mixed Review Chemical Equilibrium Answers and Bestseller Lists
- 5. Accessing Mixed Review Chemical Equilibrium Answers Free and Paid eBooks
  - Mixed Review Chemical Equilibrium Answers Public Domain eBooks
  - Mixed Review Chemical Equilibrium Answers eBook Subscription Services
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- 6. Navigating Mixed Review Chemical Equilibrium Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Mixed Review Chemical Equilibrium Answers Compatibility with Devices
  - Mixed Review Chemical Equilibrium Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mixed Review Chemical Equilibrium Answers
  - Highlighting and Note-Taking Mixed Review Chemical Equilibrium Answers
  - Interactive Elements Mixed Review Chemical Equilibrium Answers
- 8. Staying Engaged with Mixed Review Chemical Equilibrium Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mixed Review Chemical Equilibrium Answers
- 9. Balancing eBooks and Physical Books Mixed Review Chemical Equilibrium Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mixed Review Chemical Equilibrium Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Review Chemical Equilibrium Answers
  - Setting Reading Goals Mixed Review Chemical Equilibrium Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Review Chemical Equilibrium Answers
  - Fact-Checking eBook Content of Mixed Review Chemical Equilibrium Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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