Mixed Gas Laws Worksheet - Solutions

 How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

2) If 5.0 moles of O₂ and 3.0 moles of N₂ are placed in a 30.0 L tank at a temperature of 25° C, what will the pressure of the resulting mixture of gases be? 25° C = 298 K

$$P_{Tot} = P_{02} + P_{N2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0° C. By noon the temperature has risen to 45.0° C. What is the new volume of the balloon?

$$T_1 = 20.0^{\circ} C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^{\circ} C = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 L)(318 K)}{(293 K)} = 38.0 L$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in ⁶C of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol K})} = 296 \text{ K} = 23^{\circ} \text{ C}$$

6) CaCO₂ decomposes at 1200° C to form CO₂ gas and CaO. If 25 L of CO₂ are collected at 1200° C, what will the volume of this gas be after it cools to 25° C?

$$T_1 = 1200^{\circ} C = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^{\circ} C = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 L)(298 K)}{(1473 K)} = 5.1 L$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at 20.0° C is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is -20.0° C?

$$P_1$$
 = 1.00 atm, V_1 = 4.50 L, T_1 = 20.0° C = 293 K, P_2 = 0.600 atm, V_2 = ?, T_2 = -20.0° C = 253 K

$$V_2 = P_1 V_1 T_2 = (1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K}) = 6.48 \text{ L}$$

 $T_1 P_2$ (293 K)(0.600 atm)

8) There are 135 L of gas in a container at a temperature of 260° C. If the gas was cooled until the volume decreased to 75 L, what would the temperature of the gas be?

$$T_1 = 260^{\circ} C = 533 \text{ K}, V_1 = 135 \text{ L}, T_2 = 7, V_2 = 75 \text{ L}$$

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(75 \text{ L})(533 \text{ K})}{(135 \text{ L})} = 296 \text{ K} = 23^{\circ} \text{ C}$$

Mixed Gas Law Worksheet Answers

Brendan G. Carr

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