

## Mixed Nomenclature

*Identify each compound as Ionic, Covalent, or Acid, then complete the missing information.*

	Type of Compound (Ionic, Covalent, or Acid)	Name	Chemical Formula
1.	Ionic	<sup>+2</sup> Magnesium sulfide <sup>-2</sup>	MgS
2.	Covalent	Tetrasulfur octaoxide	S <sub>4</sub> O <sub>8</sub>
3.	Ionic	Magnesium chloride	MgCl <sub>2</sub>
4.	Ionic	<sup>+1</sup> Cesium carbonate <sup>-2</sup>	Cs <sub>2</sub> CO <sub>3</sub>
5.	Acid	Hydrofluoric acid	HF
6.	Covalent	Octanitrogen hexaoxide	N <sub>8</sub> O <sub>6</sub>
7.	Ionic	<sup>+2</sup> Manganese (II) chloride <sup>-1</sup>	MnCl <sub>2</sub>
8.	Acid	Phosphoric acid	H <sub>3</sub> PO <sub>4</sub>
9.	Ionic	<sup>+1</sup> Ammonium sulfate <sup>-2</sup>	(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>
10.	Ionic	Iron (III) sulfide	Fe <sub>2</sub> S <sub>3</sub>
11.	Acid	Chloric acid	HClO <sub>3</sub>
12.	Ionic	<sup>+3</sup> Aluminum chlorate <sup>-1</sup>	Al(ClO <sub>3</sub> ) <sub>3</sub>
13.	Covalent	Dinitrogen trioxide	N <sub>2</sub> O <sub>3</sub>
14.	Ionic	<sup>+1</sup> Ammonium nitrite <sup>-1</sup>	NH <sub>4</sub> NO <sub>2</sub>
15.	Ionic	Sodium nitride	Na <sub>3</sub> N
16.	Covalent	Trisulfur nonafluoride	S <sub>3</sub> F <sub>9</sub>
17.	Ionic	Copper (II) acetate	Cu(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>2</sub>
18.	Ionic	Calcium phosphate	Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>
19.	Acid	Nitrous acid	HNO <sub>2</sub>
20.	Ionic	<sup>+2</sup> Beryllium cyanide <sup>-1</sup>	Be(CN) <sub>2</sub>
21.	Covalent	Phosphorus pentabromide	PBr <sub>5</sub>
22.	Ionic	<sup>+4</sup> Tin (IV) carbonate <sup>-2</sup>	Sn(CO <sub>3</sub> ) <sub>2</sub>
23.	Acid	Hydronitric acid	H <sub>3</sub> N
24.	Ionic	Nickel (II) sulfide	NiS
25.	Covalent	Tetrachlorine heptabromide	Cl <sub>4</sub> Br <sub>7</sub>

# Mixed Nomenclature And Properties Worksheet

**R Barnett**



## **Mixed Nomenclature And Properties Worksheet:**

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*Commerce Business Daily* ,2000

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