

23. When this reaction comes to equilibrium, will the concentrations of the reactants or products be greater? Does the answer to this question depend on the initial concentrations of the reactants and products?



27. This reaction has an equilibrium constant of $K_p = 2.26 \times 10^4$ at 298 K:



Calculate K_p for each reaction and predict whether reactants or products will be favored at equilibrium:

- $CH_3OH(g) \rightleftharpoons CO(g) + 2 H_2(g) \quad 1/2.26e4 = 4.42e-5$
- $\frac{1}{2} CO(g) + H_2(g) \rightleftharpoons \frac{1}{2} CH_3OH(g) \quad (2.26e4)^{1/2} = 150.33$
- $2 CH_3OH(g) \rightleftharpoons 2 CO(g) + 4 H_2(g) \quad 1/(2.26e4)^2 = 6.65e-3$

33. Write an equilibrium expression for each chemical equation involving one or more solid or liquid reactants or products.

- $CO_3^{2-}(aq) + H_2O(l) \rightleftharpoons HCO_3^-(aq) + OH^-(aq) \quad K = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}]}$
- $2 KClO_3(s) \rightleftharpoons 2 KCl(s) + 3 O_2(g) \quad K = [O_2]^3$
- $HF(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + F^-(aq) \quad K = \frac{[H^+][F^-]}{[HF]}$
- $NH_3(aq) + H_2O(l) \rightleftharpoons NH_4^+(aq) + OH^-(aq) \quad K = \frac{[NH_4^+][OH^-]}{[NH_3]}$

35. Consider the reaction:



$$K = \frac{[CH_3OH]}{[CO][H_2]^2} = \frac{(0.185)}{(0.105)(0.105)^2} = 136$$

An equilibrium mixture of this reaction at a certain temperature has $[CO] = 0.105 M$, $[H_2] = 0.114 M$, and $[CH_3OH] = 0.185 M$. What is the value of the equilibrium constant (K_c) at this temperature?

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