Question 9. How many oxygen atoms are there in 18 g of water?

Answer: Molar mass of water is 18 g/mol.

Number of oxygen atoms is 18 g of water = 6.02 x 10²³

Question 10. Name two factors that introduce

uncertainty into measured figures.

Answer: (i) Reliability of measuring instrument.

(ii) Skill of the person making the measurement.

Question 11. State Avogadro's law.
Answer: Equal volumes of all gases under the conditions of same temperature and pressure contain the same number of molecules.

Question 12. How are 0.5 ml of NaOH differents from 0.5 M of NaOH?

Answer: 0.5 ml of NaOH means 0.5 mole (20.0 g) of NaOH,

Question 13. What is one a.m.u. or one 'u'?

Answer: 1 a.m.u. or 1 u = 1/12 th mass of an atom of carbon 12.

Question 14. What is the number of significant figures in 1.050 x 104?
Answer: Four.

0.5M of NaOH means that 0.5 mole (20.0g) of NaOH are

II. Short Answer Type Questions
Question 1. Define molality. How does molality
depend on temperature?
Answer: Molality is defined as the moles of solute per
kilogram of solvent.

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Nivaldo J. Tro

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