

CHEMICAL KINETICS

Pg-1

Rate of Reaction :

$$\text{Rate} = \frac{\text{Change in Conc or Pressure}}{\text{Time taken}}$$

$$\text{Unit} \rightarrow \text{mol L}^{-1} \text{T}^{-1} \text{ or atm T}^{-1}$$

* ROR can never be negative. (-ve)

Rate of Reaction. (A \rightarrow B)

Average Rate

$$\text{Rate} = -\frac{\Delta(A)}{\Delta t} = \frac{\Delta B}{\Delta t}$$



$$R_{\text{avg}} = \left(\frac{A_1 - A_2}{t_2 - t_1} \right)$$

Instantaneous Rate

$$\text{Rate} = -\frac{d(A)}{dt} = \frac{d(B)}{dt}$$



$$\text{Rate} = \text{slope} = \left| \frac{dy}{dx} \right|$$



$$\text{Rate (ROR)} = -\frac{d(\text{N}_2)}{dt} = -\frac{d(\text{H}_2)}{3dt} = \frac{1}{2} \frac{d(\text{NH}_3)}{dt}$$

$$\text{Rate of Disappearance} = -\frac{d(\text{N}_2)}{dt} = -\frac{d(\text{H}_2)}{dt}$$

$$\text{Rate of Appearance} = \frac{d(\text{NH}_3)}{dt} \text{ (Product)}$$

ROR & RODA does not include the St. Coeff.



$$\text{ROR} = -\frac{\text{RODA}}{a} = \frac{\text{RODB}}{b} = \frac{\text{RODC}}{c} = \frac{\text{RODD}}{d}$$

$$\text{ROR} = -\frac{1}{a} \frac{d(A)}{dt} = -\frac{1}{b} \frac{d(B)}{dt} = \frac{1}{c} \frac{d(C)}{dt} = \frac{1}{d} \frac{d(D)}{dt}$$

~ Differential Rate Law

Q. The rate of the reaction: $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$ can be written in three ways \rightarrow

$$-\frac{d(\text{N}_2\text{O}_5)}{dt} = K [\text{N}_2\text{O}_5] ; \frac{d(\text{NO}_2)}{dt} = k' [\text{N}_2\text{O}_5]$$

$$\frac{d(\text{O}_2)}{dt} = K'' [\text{N}_2\text{O}_5]$$

$$\text{Soln. ROR} = -\frac{1}{2} \frac{d(\text{N}_2\text{O}_5)}{dt} = \frac{1}{4} \frac{d(\text{NO}_2)}{dt} = \frac{d(\text{O}_2)}{dt}$$

$$= -\frac{K [\text{N}_2\text{O}_5]}{2} = \frac{K' [\text{N}_2\text{O}_5]}{4} = \frac{K'' [\text{N}_2\text{O}_5]}{1}$$

$$\text{ROR} = K' = 2K \quad K'' = \frac{K}{2}$$

$$Q = 0.3 + \log \left(-\frac{d(A)}{dt} \right) = 0.18 + \log \left(\frac{d(B)}{dt} \right)$$

write the reaction.

$$0.3 + \log \left(-\frac{d(A)}{dt} \right) = 0.18 + \log \left(\frac{d(B)}{dt} \right)$$

$$0.3 + \log \left(-\frac{d(A)}{dt} \right) = 0.18 + \log \left(\frac{d(B)}{dt} \right)$$



Molecularity :

Δ Can be integer (1, 2 or 3) but it cannot be zero, negative or fractional.

Δ For elementary reaction: Molecularity is equal to its order.

Δ Maximum value of molecularity is 3. (or order)

Δ For complex rxn molecularity of each step of mechanism is defined separately. \therefore molecularity cannot be defined for whole complex reaction at once.

Δ Molecularity of a reaction may or may not be equal to order of reaction.

Δ Molecularity can not be determined by balanced chemical rxn and can only be determined from simple reaction.

Manual Solution Of Chemical Kinetics

Alexander N. Gorban, Iliya V. Karlin



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