

10.3**THE NATURE OF SOLIDS****SECTION REVIEW****Objectives**

- Describe how the degree of organization of particles distinguishes solids from gases and liquids
- Distinguish between a crystal lattice and a unit cell
- Explain how allotropes of an element differ

Key Terms

- melting point
- crystal
- unit cell
- allotropes
- amorphous solids
- glasses

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Solids tend to be dense and ____1____. They do not flow or take the shape of their containers, like liquids do, because the particles in solids vibrate around ____2____ points. When a solid is sufficiently heated and its particles vibrate so strongly that they are no longer held in fixed positions, the solid ____3____. The ____4____ is the temperature at which a solid turns to liquid. The melting and ____5____ of a substance are at the same temperature. In general, ionic solids tend to have relatively ____6____ melting points while molecular solids tend to have relatively low melting points. Most solids are ____7____ in nature. The particles are arranged in a pattern known as a crystal ____8____. The smallest subunit of a crystal lattice is the ____9____. Some solids have a disordered array of particles and are ____10____.

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

7. _____

8. _____

9. _____

10. _____

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