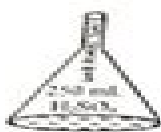


# অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটাসিয়াম কপ্লো, এখানে মোট পটাসিয়াম, আরও কপ্লোয়াম পটাসিয়ামের পটাসিয়াম এবং ময়েল টেম্পের পটাসিয়াম ল্যাবরেটরির অধ্যয়নিক সেওয়া হয়েছে। এগুলোর অনুশীলন করলে তুমি এ অধ্যয়নিক থেকে পেরোনের ল্যাবরেটরির জ্ঞানমূলক এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেলস পানিদ্রুত লবণের প্রাকৃত্য, আপনাত্মা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিভের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন ফাটল হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেলস পানিদ্রুত লবণের প্রাকৃত্য আপনাত্মা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা বৃষ্টির সাথে সাথে প্রাকৃত্য বৃষ্টি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কেলস পানি দ্রুত লবণ নিষ্কৃতি হয়। তখন প্রাকৃত্য হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পানিতে প্রকণিত করার সময় বিয়োজন প্রক্রিয়াটি আপনাত্মার  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি আপোৎপাদী হয়। এ কারণে  $32.4^\circ\text{C}$  আপনাত্মার উপরে আপনাত্মা বৃষ্টিজনিত কারণে প্রাকৃত্য হ্রাস ঘটে।

১

আমরা জানি,  
 $w = SMV$   
 $= (0.5 \times 98 \times 0.25) \text{ g}$   
 $= 12.25 \text{ g}$

অতএব, প্রকণিত  $\text{H}_2\text{SO}_4$  এর ভর  
 $= 12.25 \text{ g}$

উদীপকের A-পাত্রের অবস্থিত,  
 $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$

প্রকণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
 প্রকণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকণিত  $\text{H}_2\text{SO}_4$  এর ভর,  $w = ?$

প্রকণ-A:  $\text{H}_2\text{SO}_4$  এর মোল সংখ্যা  $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.125 \text{ mol}$

প্রকণ-B:  $\text{Na}_2\text{CO}_3$  এর মোল সংখ্যা  $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.077 \text{ mol}$

সম্প্রতি বিক্রিয়া:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol  $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে

$0.077 \text{ mol}$   $= 0.077 \text{ mol}$   $= 0.077 \text{ mol}$

অবশিষ্ট  $\text{H}_2\text{SO}_4$  এর পরিমাণ  $= (0.125 - 0.077) \text{ mol}$   
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol  $\text{H}_2\text{SO}_4$  অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2%  $\frac{W}{V}$  HCl এর নমুনা কাল করে।

- নোড কাকে বলে?
- $\text{Cl}_2$  অপোলার— ব্যাখ্যা করো।
- উদীপকের এলিভটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন। অপরদিকে, আয়নিক বন্ধনে তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর বেশি হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য 3.0 এর বেশি হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $\text{Cl}_2$  বন্ধন সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এটি প্রযোজ্য থাকে।

উদীপকে বর্ণিত এলিভটি হলো 2%  $\left(\frac{W}{V}\right)$  HCl।

2%  $\left(\frac{W}{V}\right)$  HCl এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " "  $\left(\frac{2 \times 1000}{100}\right) \text{ g}$   
 $= 20 \text{ g}$

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**M. Cusiter,V. Cusiter**



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