

Measurements Using Electrochemical Cells and Electroplating

The basic counting unit in chemistry, the mole, has a special name, *Avogadro's number*, in honor of the Italian scientist Amadeo Avogadro (1776-1856). The commonly accepted definition of Avogadro's number is the number of atoms in exactly 12 g of the isotope ^{12}C , and the quantity itself is $6.02214199(47) \times 10^{23}$.¹ A bit of information about Avogadro seems appropriate. His full name was Lorenzo Romano Amedeo Carlo Avogadro (almost a mole of letters in his name). He was a practicing lawyer until 1806 when he began his new career teaching physics and math at the University of Turin, where he was later promoted to the chair of physical chemistry. In 1811, Avogadro published a paper in the *Journal de Physique*, entitled "Essay on a Manner of Determining the Relative Masses of the Elementary Molecules of Bodies, and the Proportions in Which They Enter into These Compounds," which pretty much says it all. This paper includes the statement that has come to be regarded as Avogadro's Hypothesis:

The first hypothesis to present itself in this connection, and apparently even the only admissible one, is the supposition that the number of integral molecules in any gases is always the same for equal volumes, or always proportional to the volumes. Indeed, if we were to suppose that the number of molecules contained in a given volume were different for different gases, it would scarcely be possible to conceive that the law regulating the distance of molecules could give in all cases relations as simple as those which the facts just detailed compel us to acknowledge between the volumes and the number of molecules.

In this experiment, you will confirm Avogadro's number by conducting an electrochemical process called *electrolysis*. In electrolysis, an external power supply is used to drive an otherwise *nonspontaneous* reaction. You will use a copper strip and a zinc strip as the electrodes, placed in a beaker of sulfuric acid. You will make the cell electrolytic by using the copper strip as the anode and the zinc strip as the cathode. By determining the average current used in the reaction, along with the knowledge that all of the copper ions formed are the 2^+ cations, you will calculate the number of atoms in one mole of copper and compare this value with Avogadro's number.

OBJECTIVES

In this experiment, you will

- Prepare an electrochemical cell to oxidize a copper electrode.
- Measure the amount of copper that was deposited in the electroplating process and determine the average current used.
- Calculate a value for Avogadro's number and compare it to the accepted value.
- Calculate the value of a Faraday.

MATERIALS

Lab Quest
Current probe
Battery pack with 2 D batteries
2 wires with banana plug
3 wires with alligator clips
1 M H_2SO_4 solution
steel wool or sandpaper
balance

distilled water
copper strip (anode) (+)*
zinc strip (cathode) (-)*

* Remember the acronym "EPA"—An Electrolytic (nonspontaneous) cell has a Positive Anode.

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BM King



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Kambhoja, Dr. Priti Rajendra Kale, The past several decades have seen a tremendous change in the area of pharmaceutical analysis, mostly due to the growing complexity of the difficulties we seek to solve, the integration of interdisciplinary methods, and the development of analytical techniques. This book, *Pharmaceutical Analysis*, aims to provide a thorough review that takes into account the complex character of the state of the art studies in this area. It is intended to provide experts, researchers, and students with the fundamental abilities and information required to successfully and properly traverse the challenging field of pharmaceutical analysis. The effort was motivated by the understanding that thorough analysis procedures are essential to generating accurate and significant information. In a time where data is easily accessible and plentiful, the difficulty is not just gathering information but also arranging, analyzing, and assessing it critically in order to make significant judgments. By giving readers a comprehensive grasp of both qualitative and quantitative analytical methodologies within the context of pharmaceutical analysis, this book seeks to close the knowledge gap between theoretical principles and actual execution. Each chapter has been painstakingly written to cover a broad spectrum of subjects ranging from the sophisticated techniques used in pharmaceutical analysis to the basic concepts of analytical chemistry. Current concerns include the use of cutting-edge technology in drug research, the morality of data collecting, and the value of multidisciplinary methods, which have received special attention. This book shows how to use a variety of analytical approaches to different parts of pharmaceutical analysis via case studies and real-world examples. Putting this book together has been a joyful and difficult journey. It has included a thorough analysis of the literature, professional contacts, and the synthesis of many viewpoints. Through the joint efforts of researchers, academics, and practitioners, the information has been molded and made relevant to both present and future research pursuits. *Cracking the AP Chemistry, 2002-2003 Edition* Paul

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