



# Legacy 1 2 Hp 696cd Manual

**Lingsheng Yao**



**Legacy 1 2 Hp 696cd Manual:**

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web answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta H$  must also be halved 3 answer c explanation option c is incorrect because 566 kJ/mol is the  $\Delta H$  per mole of  $\text{O}_2$  and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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web thermochemistry practice worksheet answer key what is the heat change when 472 g of carbon reacts with excess  $\text{O}_2$  according to the following equation how much heat energy must be expelled by the body to rid the body of one pound of fat  $2\text{C}_3\text{H}_7\text{O}_2\text{(s)} + 3\text{H}_2\text{O(l)} \rightarrow 7\text{CO}_2\text{(g)} + 14\text{H}_2\text{(g)}$   $\Delta H = 7914\text{ kJ}$  is this endothermic or exothermic 2

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