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10/12

Date \_\_\_\_\_ Name \_\_\_\_\_  
Section \_\_\_\_\_ Time \_\_\_\_\_  
Instructor \_\_\_\_\_

## Pre-Lab Study Questions

1. Why are burning candles and rusting nails examples of chemical change?

Rust is oxidation

Candles are combustion  
things undergo one of the 4 criteria  
formation of gas, formation of solid, formation  
of new substances, or heat produced/absorbed.

2. What is included in a chemical equation?

a reactant and

a product

3. How does a combination reaction differ from a decomposition reaction?

Combination - two elements combine to form  
a compound.

Decomposition - the compound splits  
apart.

4. Balance each of the following reactions, and identify the type of reaction.

Unbalanced Equation	Type of Reaction
a. $2\text{Al}(s) + \text{Fe}_2\text{O}_3(s) \xrightarrow{\Delta} \text{Al}_2\text{O}_3 + \text{Fe}(l)$	single replacement
b. $2\text{KClO}_3(s) \xrightarrow{\Delta} 2\text{KCl}(s) + 3\text{O}_2(g)$	decomposition
c. $\text{I}_2(s) + \text{Cl}_2(g) \xrightarrow{\Delta} 2\text{LiCl}(s)$	combination
d. $\text{C}_2\text{H}_4(g) + \text{O}_2(g) \xrightarrow{\Delta} \text{CO}_2(g) + \text{H}_2\text{O}(g)$	double replacement
e. $2\text{CrCl}_3(aq) + 3\text{H}_2\text{S}(g) \longrightarrow \text{Cr}_2\text{S}_3(s) + 6\text{HCl}(aq)$	double replacement

5. Complete and balance each of the following reactions.

Reactants	Type of Reaction
a. $\text{K}(s) + \text{N}_2(g) \longrightarrow 2\text{KN}(g)$	Combination
b. $\text{C}_3\text{H}_{12}(g) + \text{O}_2(g) \xrightarrow{\Delta}$	Combustion
c. $\text{Al}(s) + \text{CuSO}_4(aq) \longrightarrow$	Single Replacement
d. $\text{CoCl}_2(aq) + \text{AgNO}_3(aq) \longrightarrow$	Double Replacement
e. $\text{MgCO}_3(s) \xrightarrow{\Delta} \text{MgO}(s) + \text{CO}_2(g)$	Decomposition

# Is3230 Lab Manual Questions

**University Of U. C. Davis Lab manual**



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