

- Which event must always occur for a chemical reaction to take place?
 A. formation of a precipitate
 B. formation of a gas
 C. effective collisions between reacting particles
 D. addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
 A. decreases
 B. increases
 C. remains the same
- A 5.0-gram sample of zinc and a 50.-mL 1.0 M HCl(aq) sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
 A. a zinc strip and 1.0 M HCl(aq)
 B. a zinc strip and 3.0 M HCl(aq)
 C. zinc powder and 1.0 M HCl(aq)
 D. zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 mL of 1.0 M HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
 A. 1.0 M HCl(aq) at 20.°C
 B. 1.0 M HCl(aq) at 40.°C
 C. 3.0 M HCl(aq) at 20.°C
 D. 3.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
 A. lump
 B. bar
 C. powdered
 D. sheet metal
- Given the balanced equation representing a reaction:

$$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$$

 Which set of reaction conditions produces $\text{H}_2\text{(g)}$ at the fastest rate?
 A. a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
 B. a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
 C. 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
 D. 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
 A. providing an alternate reaction pathway that has a higher activation energy
 B. providing an alternate reaction pathway that has a lower activation energy
 C. using the same reaction pathway and increasing the activation energy
 D. using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:



Which statement is true about energy in this reaction?

- The reaction is exothermic because it releases heat.
 - The reaction is exothermic because it absorbs heat.
 - The reaction is endothermic because it releases heat.
 - The reaction is endothermic because it absorbs heat.
- Which balanced equation represents an endothermic reaction?
 A. $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
 B. $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + 2\text{H}_2\text{O(g)}$
 C. $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 D. $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO(g)}$

table I
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- Given the balanced equation:



Which statement best describes this process?

- It is endothermic and entropy increases.
 - It is endothermic and entropy decreases.
 - It is exothermic and entropy increases.
 - It is exothermic and entropy decreases.
- Which change is exothermic?
 A. freezing of water
 B. melting of iron
 C. vaporization of ethanol
 D. sublimation of iodine
 - Which reaction releases the greatest amount of energy per 2 moles of product?
 A. $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
 B. $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$
 C. $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$
 D. $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$

look up values
 on table I

- Given the balanced equation representing a reaction at 100.3 kPa and 298 K:



Which statement is true about this reaction?

- It is exothermic and ΔH equals -91.8 kJ.
 - It is exothermic and ΔH equals +91.8 kJ.
 - It is endothermic and ΔH equals -91.8 kJ.
 - It is endothermic and ΔH equals +91.8 kJ.
- Which expression represents the ΔH for a chemical reaction in terms of the potential energy, PE, of its products and reactants?
 A. PE of products + PE of reactants
 B. PE of products - PE of reactants
 C. PE of products \times PE of reactants
 D. PE of products \div PE of reactants

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İstanbul da cambridge english preliminary pet dersi verenler - May 12 2023

web cambridge english preliminary pet okul Öncesi İlköğretim takviye lise takviye Üniversite takviye sınav hazırlık yabancı dil bilgisayar müzik spor sanat dans kişisel gelişim direksiyon Özel eğitim oyun ve hobi

pet walker and care student jobs istanbul erasmusu com - Dec 07 2022

web student jobs in istanbul need to leave the pet because of vacation business or holidays for few days and looking for a clean safety shelter for all pets as hamsters guineapigs birds dogs etc except big dogs and cats because have 2 own dogs or someone to come to your place and check your pet don t hesitate to

should college students have pets ist 110 introduction to - Aug 03 2022

web oct 10 2021 although your post relates directly to college students having a dog i believe pets in general are a great

support system for college students with the stress and anxiety that comes from meeting deadlines and social needs having a friend no matter what kind of animal can help

8 benefits of classroom pets with infographic pet keen - Aug 15 2023

web aug 21 2023 having a pet in the classroom will give you the opportunity to teach a variety of lessons based on the animal itself a classroom animal can provide an interactive way to engage the children with their studies and you can use science math and even english assignments based on the pet