

## PHYSICAL KINETIC MOLECULAR THEORY

### Learning Objectives

- 1) Identify theoretical differences in how gases participate in the kinetic theory of gases and how molecules in liquids and solids participate in the kinetic theory of gases and molecular theory.
- 2) Identify how temperature affects molecular motion.
- 3) Explain how pressure affects molecular motion and how it is related to molecular theory.

### Module 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in gas, liquid, and solid phases.



Figure 1: Representation of particles in different phases of matter.

Understanding the differences in how particles participate in the kinetic theory of gases and how molecules in liquids and solids participate in the kinetic theory of gases is essential for understanding the behavior of matter in different phases.

### Key Concepts

1. **Gas:** Gas particles are widely spaced and move rapidly in all directions. They have high kinetic energy and are not held together by strong intermolecular forces.
2. **Liquid:** Liquid particles are in contact but have some free space. They move more slowly than gas particles and are held together by moderate intermolecular forces.
3. **Solid:** Solid particles are in a highly ordered, repeating lattice structure. They have low kinetic energy and are held together by strong intermolecular forces.

Table 1: Comparison of the Kinetic Theory of Gases.

	Gas	Liquid	Solid
<b>Particle Arrangement</b>	Particles are widely spaced and move rapidly in all directions. They have high kinetic energy and are not held together by strong intermolecular forces.	Particles are in contact but have some free space. They move more slowly than gas particles and are held together by moderate intermolecular forces.	Particles are in a highly ordered, repeating lattice structure. They have low kinetic energy and are held together by strong intermolecular forces.
<b>Particle Motion</b>	Particles move rapidly in all directions. They have high kinetic energy and are not held together by strong intermolecular forces.	Particles move more slowly than gas particles and are held together by moderate intermolecular forces.	Particles move very slowly and are held together by strong intermolecular forces.

# Lw Pogil Kinetic Molecular Theory Key

**S Marginson**



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