

Chemistry Paper 1 revision

ELECTROLYSIS:

- *Separating elements in insoluble ionic compounds*

Equipment-

- *Beaker containing an electrolyte (liquid/solution containing an ionic compound) → ions are free to move*
- *Electrodes, which are solid conductors (usually metal or carbon), positive electrode (anode), negative electrode (cathode)*
- *Wire, used to join the electrodes so the electrons can flow between the electrodes*
- *Add a power supply to the wire, such as a battery to drive the flow of electrons*

Electrolysis- 'splitting up with electricity', by using the battery to pass an electric current through the electrolyte.

Negatively charged ions go to the opposite charge such as the positively charged ions do, for example:

- *Negatively charged ions → Anode (positive electrolyte)*
- *Positively charged ions → Cathode (negative electrolyte)*

Once these ions go to the oppositely charged electrolyte it becomes discharged which means that it becomes a neutral atom such as Br^- would become Br (negatively charged bromine ion becomes bromine)

Electrons are reduced at a cathode, and are increased at an anode

PART TWO

Electrolysis to extract reactive metals from their oxides:

- Examples of metal oxides would be aluminium oxide and copper oxide, the goal is to separate the oxygen molecules from the metal, (remove the oxygen)
- One way this can be achieved is through reduction with carbon, where the carbon displaces the metal, and the equation would be
metal oxide + carbon → pure metal + carbon dioxide. However, this only works for metals that are less reactive than carbon.

REACTIVITY SERIES-

Potassium

Sodium

Lithium

Calcium

Magnesium

CARBON

Zinc

Iron

Hydrogen

Copper

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Rachel Sandford



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