# Student Solution Manual

Student Solution Manual: Skoog et al., Fundamentals of Analytical Chemistry, 10e, © 2022, 978-0-357-45055-0, Chapter 2: Calculations Used in Analytical Chemistry

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

## Chapter 2

2-1. Define

#### Answers:

(a) molar mass.

The molar mass is the mass in grams of one mole of a chemical species.

(c) millimolar mass.

The millimolar mass is the mass in grams of one millimole of a chemical species.

Give two examples of units derived from the fundamental base SI units.

#### Solution:

The liter: 
$$1 L = \frac{1000 \text{ mL}}{1 L} \times \frac{1 \text{ cm}^3}{\text{mL}} \times \left(\frac{\text{m}}{100 \text{ cm}}\right)^3 = 10^{-3} \text{ m}^3$$

Molar concentration: 
$$1 M = \frac{1 \text{ mol}}{L} \times \frac{L}{10^{-3} \text{ m}^3} = \frac{1 \text{ mol}}{10^{-3} \text{ m}^3}$$

Simplify the following quantities using a unit with an appropriate prefix:

#### Solutions:

(a) 5.8 × 10<sup>8</sup> Hz.

$$5.8 \times 10^8 \text{ Hz} \times \frac{\text{MHz}}{10^8 \text{ Hz}} = 580 \text{ MHz}$$

(c) 9.31 × 10<sup>3</sup> μmol.

$$9.31 \times 10^7 \text{ } \mu \text{mol} \times \frac{\text{mol}}{10^6 \text{ } \mu \text{mol}} = 93.1 \text{ mol}$$

(e) 3.96×10° nm.

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**RJ Shavelson** 

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