

EQUILIBRIUM TEST

MULTIPLE CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. It is so very large, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen trioxide, N_2O_3 , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



At a given temperature, the equilibrium constant for this reaction is 0.50. If K_c and the initial concentration of N_2O_3 at equilibrium is 0.5000 M, at this temperature, the initial concentration of NO_2 at equilibrium is

- A. 0.5000 M
- B. 1.000 M
- C. 1.50 M
- D. 2.00 M

10) Chloroacetic acid, ClCH_2COOH , reacts with water, H_2O , to form methanoic acid, H_2COOH , H_2COO^- , chloroacetate ions, $\text{ClCH}_2\text{COO}^-$ and hydrochloric acid, HCl . At equilibrium, the concentration of H_2COOH is 0.0010 M, the concentration of H_2COO^- is 0.0010 M, the concentration of ClCH_2COOH is 0.0010 M, the concentration of $\text{ClCH}_2\text{COO}^-$ is 0.0010 M, the concentration of HCl is 0.0010 M, and the concentration of H_2O is 55.5 M. The equilibrium constant is



11) Suppose that a mixture of 1.0 mol chloroacetic acid is placed in a closed and held overnight and that chloroacetic acid reaches equilibrium. In the morning, after changing the volume of the system, it is found that the following change has taken place: the volume is not affected by the concentration of H_3O^+ ions?

- A. Consider that solution.
- B. Double the moles ClCH_2COOH gas.
- C. Add a few drops of perchloric acid solution.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Multiple Choice Concept Questions On Equilibrium

Greg Jacobs



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