

## Stoichiometry: Multiple Choice

1. The limiting reactant of a chemical reaction is

- A) the reactant of a chemical process that is not consumed completely
- B) the product of a chemical process that is consumed completely
- C) the product of a chemical process that is not consumed completely
- D) the reactant of a chemical process that is consumed completely
- E) none of the above

ANS: D

1. Consider the following reaction;  $\text{Ca}(\text{OH})_{2(aq)} + \text{Al}_2(\text{SO}_4)_{3(aq)} \rightarrow \text{CaSO}_{4(s)} + \text{Al}(\text{OH})_{3(s)}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

- A) 1:1
- B) 2:1
- C) 2:3
- D) 1:2
- E) 3:1

ANS: E

1. Sodium phosphate,  $\text{Na}_3\text{PO}_4$ , can be prepared by combining sodium hydroxide and hydrogen phosphate as shown in the following balanced chemical equation:



If 2.5 mol of  $\text{H}_2\text{PO}_4$  are reacted, the number of moles of water that will be produced is

- A) 1.0
- B) 2.5
- C) 3.0
- D) 5.0
- E) 7.5

ANS: E

1. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:



The number of moles of  $\text{KClO}_3$  required to produce 5.0 mol of oxygen is

- A) 1.0
- B) 2.0
- C) 3.0
- D) 3.3
- E) 5.0

ANS: D

1. Determine the number of grams of  $\text{CO}_2$  that can be produced from 48.0 g of the hydrocarbon  $\text{C}_3\text{H}_8$ , according to the following equation:



- A) 37.6 g
- B) 50.2 g
- C) 87.8 g
- D) 151 g
- E) 188.4 g


1. The limiting reactant in a chemical reaction is the \_\_\_\_\_ that is \_\_\_\_\_.

- A) Reactant, in excess
- B) Product, left over
- C) Reactant, used up
- D) Product, used up

ANS: C

# Multiple Choice Questions On Chemical Stoichiometry

**Richard H. Langley**



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### **Table of Contents Multiple Choice Questions On Chemical Stoichiometry**

1. Understanding the eBook Multiple Choice Questions On Chemical Stoichiometry
  - The Rise of Digital Reading Multiple Choice Questions On Chemical Stoichiometry
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Chemical Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions On Chemical Stoichiometry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Chemical Stoichiometry
  - Personalized Recommendations
  - Multiple Choice Questions On Chemical Stoichiometry User Reviews and Ratings
  - Multiple Choice Questions On Chemical Stoichiometry and Bestseller Lists



5. Accessing Multiple Choice Questions On Chemical Stoichiometry Free and Paid eBooks
  - Multiple Choice Questions On Chemical Stoichiometry Public Domain eBooks
  - Multiple Choice Questions On Chemical Stoichiometry eBook Subscription Services
  - Multiple Choice Questions On Chemical Stoichiometry Budget-Friendly Options
6. Navigating Multiple Choice Questions On Chemical Stoichiometry eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions On Chemical Stoichiometry Compatibility with Devices
  - Multiple Choice Questions On Chemical Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Chemical Stoichiometry
  - Highlighting and Note-Taking Multiple Choice Questions On Chemical Stoichiometry
  - Interactive Elements Multiple Choice Questions On Chemical Stoichiometry
8. Staying Engaged with Multiple Choice Questions On Chemical Stoichiometry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions On Chemical Stoichiometry
9. Balancing eBooks and Physical Books Multiple Choice Questions On Chemical Stoichiometry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions On Chemical Stoichiometry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions On Chemical Stoichiometry
  - Setting Reading Goals Multiple Choice Questions On Chemical Stoichiometry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions On Chemical Stoichiometry
  - Fact-Checking eBook Content of Multiple Choice Questions On Chemical Stoichiometry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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