

## SOME BASIC CONCEPTS OF PHYSICAL CHEMISTRY

### **SOME BASIC CONCEPTS OF PHYSICAL CHEMISTRY**

Measurement in chemical (Precision, significant figures, SI units, Dimensional analysis), Laws of chemical combination, Atomic Mass, Molecular Mass, molar concept, Molar Mass, determination of Molecular formula, Chemical equation, stoichiometry of Chemical reactions.

### **INITIAL STEP EXERCISE**

#### **Multiple Choice Questions**

1. 1 g of the carbonate of a metal was dissolved in 25 ml of  $\text{HCl}$ . The resulting liquid required 8 ml of  $\text{N-NaOH}$  for neutralisation. The eq. wt. of the metal carbonate is  
 (a) 30 (b) 20  
 (c) 25 (d) None
2. 8 ml of  $\text{N-HCl}$ , 25 ml of  $\text{Na}_2\text{B}_4\text{O}_7$  and 20 ml of  $\text{NaOH} = \text{HNO}_3$  are mixed together and the mixture made to 1 litre.  
 (a) The normality of the resulting solution is  
 (a)  $\text{N/5}$  (b)  $\text{N/10}$   
 (c)  $\text{N/20}$  (d)  $\text{N/40}$   
 (b) The vol. of pure  $\text{NaOH}$  required to neutralise the above solution is  
 (a) 50 g (b) 5 g  
 (c) 1 g (d) 2.5 g
3. If 0.1 mol of  $\text{BaCl}_2$  is mixed with 0.20 mol of  $\text{Na}_3\text{PO}_4$ , the maximum amount of  $\text{Ba}_3(\text{PO}_4)_2$  that can be formed is  
 (a) 0.10 mol (b) 0.20 mol  
 (c) 0.25 mol (d) 0.15 mol
4. The equivalent mass of  $\text{MnO}_2$  is half of its molar mass when it is converted to  
 (a)  $\text{Mn}_2\text{O}_3$  (b)  $\text{MnCl}_2$   
 (c)  $\text{MnCl}_2^-$  (d)  $\text{MnCl}_2^{2+}$
5. The nitrate nitride can be converted into ammonium ion. The equivalent mass of  $\text{NO}_2^-$  ion in this reaction would be  
 (a) 4.20 g (b) 7.20 g  
 (c) 10.2 g (d) 22.8 g

# Multiple Choice Questions On Inorganic Chemistry

**Margaret M. L. Oblitas**



## Multiple Choice Questions On Inorganic Chemistry:

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## **Table of Contents Multiple Choice Questions On Inorganic Chemistry**

1. Understanding the eBook Multiple Choice Questions On Inorganic Chemistry
  - The Rise of Digital Reading Multiple Choice Questions On Inorganic Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Inorganic Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions On Inorganic Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Inorganic Chemistry
  - Personalized Recommendations
  - Multiple Choice Questions On Inorganic Chemistry User Reviews and Ratings
  - Multiple Choice Questions On Inorganic Chemistry and Bestseller Lists
5. Accessing Multiple Choice Questions On Inorganic Chemistry Free and Paid eBooks

- Multiple Choice Questions On Inorganic Chemistry Public Domain eBooks
- Multiple Choice Questions On Inorganic Chemistry eBook Subscription Services
- Multiple Choice Questions On Inorganic Chemistry Budget-Friendly Options
- 6. Navigating Multiple Choice Questions On Inorganic Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions On Inorganic Chemistry Compatibility with Devices
  - Multiple Choice Questions On Inorganic Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Inorganic Chemistry
  - Highlighting and Note-Taking Multiple Choice Questions On Inorganic Chemistry
  - Interactive Elements Multiple Choice Questions On Inorganic Chemistry
- 8. Staying Engaged with Multiple Choice Questions On Inorganic Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions On Inorganic Chemistry
- 9. Balancing eBooks and Physical Books Multiple Choice Questions On Inorganic Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions On Inorganic Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Questions On Inorganic Chemistry
  - Setting Reading Goals Multiple Choice Questions On Inorganic Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Questions On Inorganic Chemistry
  - Fact-Checking eBook Content of Multiple Choice Questions On Inorganic Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development



- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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