

# EQUILIBRIUM MCQ [59 marks]

1. What effect does increasing both pressure and temperature have on the equilibrium constant,  $K_c$ ? [1 mark]



- ☒ A. Decreases  
☐ B. Increases  
☐ C. Remains constant  
☐ D. Cannot be predicted as effects are opposite

2.  $K_c$  for  $2\text{N}_2\text{O} (\text{g}) \rightleftharpoons 2\text{N}_2 (\text{g}) + \text{O}_2 (\text{g})$  is  $7.3 \times 10^{34}$ . [1 mark]

What is  $K_c$  for the following reaction, at the same temperature?



- A.  $7.3 \times 10^{34}$   
☒ B.  $\frac{1}{\sqrt{7.3 \times 10^{34}}}$   
C.  $\frac{2}{7.3 \times 10^{34}}$   
D.  $\frac{1}{2 \times 7.3 \times 10^{34}}$

3. What is the equilibrium constant expression for the following equation? [1 mark]



- A.  $\frac{2[\text{NO}_2\text{F}]}{2[\text{NO}_2] + [\text{F}_2]}$  ☒  
B.  $\frac{2[\text{NO}_2\text{F}]}{2[\text{NO}_2][\text{F}_2]}$  ☒  
C.  $\frac{[\text{NO}_2]^2[\text{F}_2]}{[\text{NO}_2\text{F}]^2}$  ☒  
☒ D.  $\frac{[\text{NO}_2\text{F}]^2}{[\text{NO}_2]^2[\text{F}_2]}$

# Mcq On Chemical Equilibrium

**Marco Cascella**



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