

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- 1) Students will understand differences in how gases, liquids, and solids behave, and how these differences are explained by the postulates of the kinetic molecular theory.
- 2) Students will understand the relationship between temperature and kinetic energy.
- 3) Students will understand the relationship between pressure and molecular collisions.

Module 1: Representation of States in Different Phases

The following figure illustrates the representative particles for each phase of matter: solid, liquid, and gas.

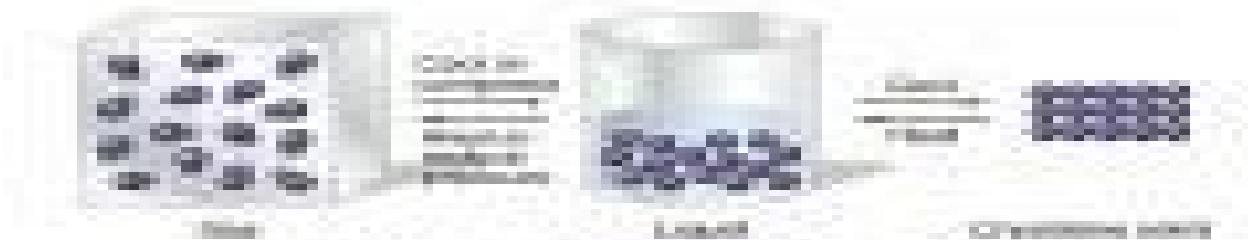


Figure 1: Representative particles in different states of matter.

Students will understand the differences in the behavior of particles in different states of matter. Students will understand the relationship between temperature and kinetic energy.

Key Concepts

1. Matter can be in three states: solid, liquid, and gas. The particles in each state have different arrangements and different types of motion. The particles in a solid are closely packed and have a fixed shape. The particles in a liquid are closely packed but have a variable shape. The particles in a gas are widely spaced and have a variable shape.

Table 1: Properties of Matter in Different States

	SOLID	LIQUID	GAS
Arrangement of particles	Particles are closely packed in a regular arrangement. They have a fixed shape.	Particles are closely packed in a disordered arrangement. They have a variable shape.	Particles are widely spaced in a disordered arrangement. They have a variable shape.
Distance between particles	Distance between particles is very small. They are in direct contact with each other.	Distance between particles is small. They are in direct contact with each other.	Distance between particles is large. They are not in direct contact with each other.
Force of attraction	Force of attraction between particles is very strong. They are held together tightly.	Force of attraction between particles is moderate. They are held together loosely.	Force of attraction between particles is very weak. They are not held together.
Compressibility	Particles are closely packed. They cannot be compressed further.	Particles are closely packed. They can be compressed slightly.	Particles are widely spaced. They can be compressed easily.

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