

CHAPTER **11** STUDY GUIDE

Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write *true* or *false*.

- _____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
- _____ 2. Stoichiometry is based on the law of conservation of mass.
- _____ 3. In any chemical reaction, the mass of the products is less than the mass of the reactants.
- _____ 4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
- _____ 5. The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.

methanol + oxygen \rightarrow carbon dioxide + water



Substance	Molar Mass (g/mol)	Number of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05			
7. Oxygen gas	32.00			
8. Carbon dioxide	44.01			
9. Water	18.02			

10. What are the reactants? _____
11. What are the products? _____
12. What is the total mass of the reactants? _____
13. What is the total mass of the products? _____
14. How do the total masses of the reactants and products compare? _____

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Bruce Reinhart



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Table of Contents Mcgraw Hill Business Management Study Guide Answers

1. Understanding the eBook Mcgraw Hill Business Management Study Guide Answers
 - The Rise of Digital Reading Mcgraw Hill Business Management Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Business Management Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Business Management Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Business Management Study Guide Answers
 - Personalized Recommendations
 - Mcgraw Hill Business Management Study Guide Answers User Reviews and Ratings
 - Mcgraw Hill Business Management Study Guide Answers and Bestseller Lists
5. Accessing Mcgraw Hill Business Management Study Guide Answers Free and Paid eBooks
 - Mcgraw Hill Business Management Study Guide Answers Public Domain eBooks
 - Mcgraw Hill Business Management Study Guide Answers eBook Subscription Services
 - Mcgraw Hill Business Management Study Guide Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Business Management Study Guide Answers Compatibility with Devices
 - Mcgraw Hill Business Management Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Business Management Study Guide Answers
 - Highlighting and Note-Taking Mcgraw Hill Business Management Study Guide Answers
 - Interactive Elements Mcgraw Hill Business Management Study Guide Answers
8. Staying Engaged with Mcgraw Hill Business Management Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Business Management Study Guide Answers
9. Balancing eBooks and Physical Books Mcgraw Hill Business Management Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Business Management Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Business Management Study Guide Answers
 - Setting Reading Goals Mcgraw Hill Business Management Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Business Management Study Guide Answers
 - Fact-Checking eBook Content of Mcgraw Hill Business Management Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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