- 51, 4.08 g of a mixture of BaO and an unknown carbonate MCO₃ was heated strongly. The residue weighed 3.64 g. This was dissolved in 100ml of 1 N HCL. The excess acid required 16 ml of 2.5 N NaOH solution for complete neutralization. Identify the metal M. (At. Wt. H=1, C=12, O=16, Cl=35.5, Ba=138)
 - a) Ca b) Mg c) Na d) Cl
- 52. Study the following table

Compound (molecular weight)	Weight of compound (in g) taken
I. CO, (44)	4.4
II. NO ₂ (46)	2.3
III. H ₂ O ₂ (34)	6.8
IV. SO ₂ (64)	1.6

Which two compound have least weight of oxygen?(molecular weights of compound are given in brackets)

a) II and IV
 c) I and II

- b) Land III
- d) Ill and IV
- 53. A gaseous mixture contains oxygen and nitrogen in the ratio of 1 : 4 by weight. Therefore the ratio of their number of molecules is
 - a) 1:4 c) 7:32
- b) 1:8d) 3:16
- 54. If 0.50 mole of BaCl₂ is mixed with 0.20 mole of Na₂PO₄, the maximum number of moles of Ba₂(PO₄)₂ that can be formed is
 - a) 0.70

- b) 0.50 d) 0.10
- c) 0.20 d
- Which of the following expressions correctly represents the relationship between the average molar kinetic energy, KE, of CO and

N, molecules at the same temperature?

- a) KEco≤KE_{N2}
- b) KEco> KEN;
- c) KEco=KEs,
- d) cannot be predicted unless volumes of the gases are given

- The ratio of the rate of diffusion of helium and methane under identical condition of pressure and temperature will be
 - a) 4 c) 2

- b) 0.2 d) 0.5
- An element, X has the following isotopic composition:

200X:90% 100X:8.0% 202X:2.0%

The weighted average atomic mass of the naturally occurring element X is closest to

a) 200u c) 202u

- b) 201ud) 199u
- 58. The bohr orbit radius for the hydrogen atom (n = 1) is approximately 0.530A° the radius for the first excited state (n=2) orbit is (in A°) a) 0.13 b) 1.06
 - a) 0.13 c) 4.77
- d) 2.17
- 59. Correct set of four quantum number for the valence (outermost) electron of rubidium (Z= 37) is:
 - a) 5,0,0,+1/2 c) 5 1 1 +1/2
- b) 5,1,0,+1/2 d) 6,0,0,+1/2
- c) 5,1,1,+1/2 d The correct order of
- 60. The correct order of C—O bond length among CO,CO₂², CO₂ is
 - a) CO < CO₂ < Co₂
 - b) CO2 < CO2 < CO
 - c) CO < CO₂ < Co₂²
 - d) CO, < CO; < CO
- 61. The bond length of HCL molecule is 1.275Å and its dipole moment is 1.03D, the ionic character of the molecule (in percent) (charge of the electron = 4.8 × 10⁻¹⁰ esu) is

 a) 100

 b) 67.3
 - a) 100 c) 33.66
- d) 16.83
- Match the list-I (Equations) with list-II (type of processes) and select the correct option.
 List I Equation List II
 - Non spontaneous (1) Kp≥O
- Type of Processes
 (i) Non spontaneous
- $(2)\Delta G^{\circ} \leq RT \ln Q$
- (ii) Equilibrium

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G Thomas

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