



(a) i) Suggest a suitable reagent that can be used as **solid A**. (1 mark)

ii) Name liquids **C** and **H**. (2 marks)

C

H

iii) Write a balanced chemical equation for the reaction in conical flask **B**. (1 mark)

iv) Explain why **solid U** collects further away from aluminium metal. (1 mark)

(b) During a class experiment, chlorine gas was bubbled into a solution of potassium iodide.
i) State the observation made. (1 mark)

ii) Write the ionic equation for the reaction that took place. (1 mark)

(c) Write a balanced chemical equation for the reaction between hot concentrated sodium hydroxide and chlorine gas. (1 mark)

(d) Explain the difference in bleaching by chlorine and bleaching by sulphur (IV) oxide gas. (2 marks)

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