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Name		

Acid-Base Test Review

Directions: On a separate piece of paper, answer all questions in complete sentences or show all your work for full credit.

- 1. What is an acid?
 - What are the properties of acids
 - Differentiate between strong acids and weak acids, in terms of pH and [H+]
- 2. What is a base?
 - a. What are the properties of bases
 - Differentiate between strong bases and weak bases, in terms of pH and [OH-]
- 3. Explain what pH measures, and what it tells us about acids and bases.
- 4. What is the Kw? What is the equation for Kw?
- Create a table showing the 3 acid-base definition.
- 6. Define and provide an example of the following:
 - a. Monoprotic -
 - b. Diprotic -
 - c. Triprotic -
- 7. What are conjugate acids and bases? Provide an example.
- 8. How is pH calculated?
- 9. Summarize naming rules for acids
- 10. Identify each property as applying to an acid, a base, or both.
 - a. bitter taste –
 - b. electrolyte -
 - e. indicator color change -
 - d. sour taste -
- 11. Write the formula for each acid or base
 - a. barium hydroxide –
 - b. hydrobromic acid -
 - c. rubidium hydroxide -
- 12. Name or give the molecular formula for each of the following acid or base.
 - a. Acetic acid-
 - b. Phosphorous acid-
 - c. HCLO_{4*}
 - d. H₃PO₄-
 - e. sulfuric acid-
 - f. HCLO-
 - g. HF-
 - h. HClO₃ -
 - H₂CO₃ –
 - i. Al(OH)s -
- 13. Define acidic solution -
- 14. Define basic solution -
- 15. Define pH -
- Explain how you would find the pH of a solution if you had the OH concentration.

Key Acid Base Study Guide

Steven S. Zumdahl, Paul B. Kelter

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