

Section C

Question Number	Acceptable Answers	Mark
24(a)(i)	$(K_c =) \frac{[\text{CH}_3\text{COOCH}_2\text{CH}_3][\text{H}_2\text{O}]}{[\text{CH}_3\text{COOH}][\text{CH}_3\text{CH}_2\text{OH}]}$ ALLOW C₂H₅OH for ethanol ALLOW CH₃CO₂H for ethanoic acid ALLOW CH₃CO₂CH₂CH₃ / CH₃CO₂C₂H₅ / CH₃COOC₂H₅ for ethyl ethanoate I IGNORE state symbols, even if incorrect	(1)
Question Number	Acceptable Answers	Mark
24(a)(ii)	Stand alone marks the enthalpy change is (very) small/close to zero OR reaction is slightly exothermic (1) therefore, (the magnitude of) $\Delta S_{\text{surroundings}}$ (= $-\Delta H/T$) changes very little (1) I IGNORE $\Delta S_{\text{surroundings}}$ is positive/ small/ less/ decreases ΔS_{total} / K_c changes very little (provided there is no change of state) (1) Ignore references to ΔS_{system}	(3)

January 2015 C3 Ial

Georges Romme

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Georges Romme.

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