

Literacy Content Specialty Test Study Guide

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reaction a redox reaction explain your answer $2\text{NaCl (aq)} + \text{Pb (s)} \rightarrow 2\text{NaNO}_3\text{ (aq)} + \text{PbCl}_2\text{ (s)}$

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b Ba^{2+} c Br_2 d ca e Ga^{3+} f h 2g h

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na 2 so 4 b mno 4 c KClO_3 d CH_3OH 2 identify the oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs when an atom's oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom's oxidation state becomes less positive indicating that electrons have been gained

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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 what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
 a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + 2\text{I}^-$ 5 $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation
 reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation
 state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the
 reducing agent also 2 $\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr 0 to Sr^{2+} oxidized reducing agent O_2 0 to O^{2-} reduced ox ag 2 $\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li 0 to Li^{+} oxidized red ag
 S 0 to S^{2-} reduced ox ag 3 2 $\text{Cs} + \text{Br}_2$

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