

12. Molecular Geometry and Bonding Theories



Class: Principles and Mechanisms, Semester IV, 2019

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Original Content	
<p>Learning Object: To understand hybridization and how it gives molecular geometry.</p> <p>Consider the two atomic orbitals (represented in Fig. 12.1) participating in sp³ hybridization. The two orbitals have two phases (positive and negative). When hybridized, four sp³ hybrid orbitals are formed. Each orbital has a positive or negative phase. The four sp³ hybrid orbitals are arranged in a tetrahedral geometry. The resulting orbitals are called sp³ hybrid orbitals.</p> <p>Hybridization of the four</p> 	<p>Hybridization of sp³</p> 
<p>Explain this and the corresponding steps and their progress.</p>	
<p>Step 1:</p> <p>What type of atomic orbitals are used in sp³ hybridization?</p>	
<p>Ans 1.1 Specify the electron configuration of carbon.</p> <p>What are the orbitals?</p>	
<p>Answer: Carbon-12 has 6 electrons. The configuration is 1s² 2s² 2p².</p>	
<p>Step 2:</p> <p>What type of atomic orbitals are used in sp³ hybridization?</p>	
<p>Ans 2.1 What are the orbitals? Specify the electron configuration of carbon.</p> <p>What are the orbitals?</p>	
<p>Answer: Carbon-12 has 6 electrons. The configuration is 1s² 2s² 2p².</p>	

Masteringchemistry Answer Key

Barbara Herlihy



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