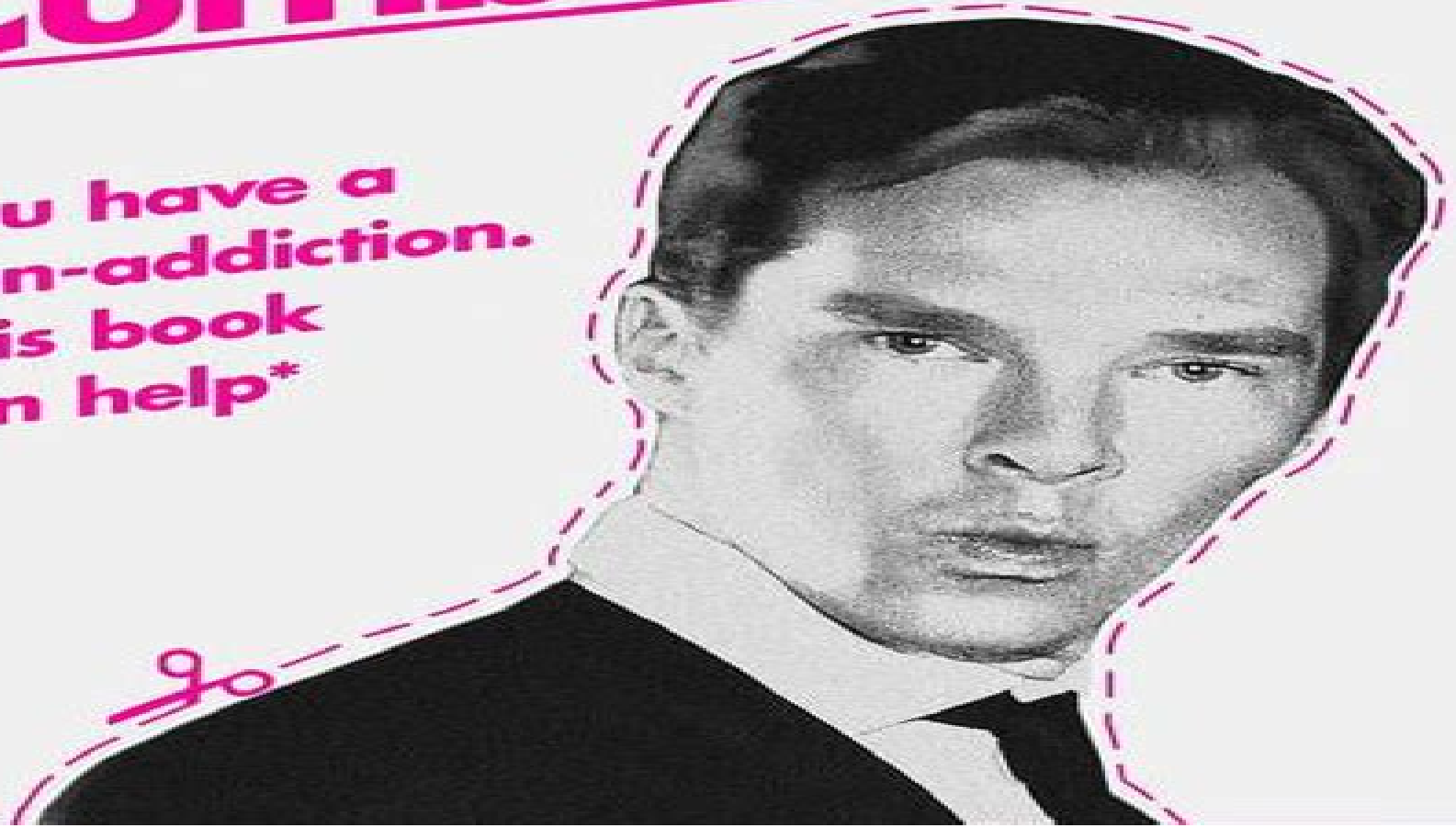


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with 1200 ml of 2M HNO_3 if the final density of solution is same as pure water then molarity of the iron in resulting solution which decides the nature of the above solution is 1 0.25M 2 0.5M 3 0.888M 4 none of

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$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$m_{mix} = \frac{m_1 \cdot V_1 + m_2 \cdot V_2}{V_1 + V_2}$$
$$1.5 \times 480 + 1.2 \times 520 = 480 + 520 = 1000$$
$$m_{mix} = \frac{1344}{1000} = 1.344 \text{ M}$$
 with increase in temperature which of these changes a molality b fraction of solute present in water

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