

2. Sodium thiosulfate solution,  $\text{Na}_2\text{S}_2\text{O}_3(\text{aq})$ , and hydrochloric acid,  $\text{HCl}(\text{aq})$ , react to produce solid sulfur as in the equation below.



The following results to determine the initial rate were obtained:

Experiment	$[\text{S}_2\text{O}_3^{2-}(\text{aq})] / \text{mol dm}^{-3}$	$[\text{H}^+(\text{aq})] / \text{mol dm}^{-3}$	Initial rate / $\text{mol dm}^{-3} \text{s}^{-1}$
1	0.200	2.00	0.036
2	0.200	1.00	0.036
3	0.100	1.00	0.018

- (a) Deduce, with a reason, the order of reaction with respect to each reactant.

[2]

- (b) State the rate expression for this reaction.

[1]

- (c) Determine the value of the rate constant,  $k$ , and state its units.

[2]

# N14 4 Mathl Hp2 Tz1

**Ying Liu**



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## **Table of Contents N14 4 Mathl Hp2 Tz1**

1. Understanding the eBook N14 4 Mathl Hp2 Tz1
  - The Rise of Digital Reading N14 4 Mathl Hp2 Tz1
  - Advantages of eBooks Over Traditional Books
2. Identifying N14 4 Mathl Hp2 Tz1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an N14 4 Mathl Hp2 Tz1
  - User-Friendly Interface
4. Exploring eBook Recommendations from N14 4 Mathl Hp2 Tz1
  - Personalized Recommendations
  - N14 4 Mathl Hp2 Tz1 User Reviews and Ratings
  - N14 4 Mathl Hp2 Tz1 and Bestseller Lists
5. Accessing N14 4 Mathl Hp2 Tz1 Free and Paid eBooks
  - N14 4 Mathl Hp2 Tz1 Public Domain eBooks
  - N14 4 Mathl Hp2 Tz1 eBook Subscription Services
  - N14 4 Mathl Hp2 Tz1 Budget-Friendly Options

6. Navigating N14 4 Mathl Hp2 Tz1 eBook Formats
  - ePub, PDF, MOBI, and More
  - N14 4 Mathl Hp2 Tz1 Compatibility with Devices
  - N14 4 Mathl Hp2 Tz1 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of N14 4 Mathl Hp2 Tz1
  - Highlighting and Note-Taking N14 4 Mathl Hp2 Tz1
  - Interactive Elements N14 4 Mathl Hp2 Tz1
8. Staying Engaged with N14 4 Mathl Hp2 Tz1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers N14 4 Mathl Hp2 Tz1
9. Balancing eBooks and Physical Books N14 4 Mathl Hp2 Tz1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection N14 4 Mathl Hp2 Tz1
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine N14 4 Mathl Hp2 Tz1
  - Setting Reading Goals N14 4 Mathl Hp2 Tz1
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of N14 4 Mathl Hp2 Tz1
  - Fact-Checking eBook Content of N14 4 Mathl Hp2 Tz1
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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