

## IB SL and HL Topic 1 to 5 Chemistry MCQ 40 Marks

1. What amount of oxygen,  $O_2$ , (in moles) contains  $1.8 \times 10^{22}$  molecules?

- A. 0.0030
- B. 0.030
- C. 0.30
- D. 3.0

(Total 1 mark)

2. Which compound has the empirical formula with the greatest mass?

- A.  $C_2H_6$
- B.  $C_4H_{10}$
- C.  $C_3H_{10}$
- D.  $C_6H_6$

(Total 1 mark)

3.  $\_\_ C_2H_2(g) + \_\_ O_2(g) \rightarrow \_\_ CO_2(g) + \_\_ H_2O(g)$

When the equation above is balanced, what is the coefficient for oxygen?

- A. 2
- B. 3
- C. 4
- D. 5

(Total 1 mark)

4.  $3.0 \text{ dm}^3$  of sulfur dioxide is reacted with  $2.0 \text{ dm}^3$  of oxygen according to the equation below.



What volume of sulfur trioxide (in  $\text{dm}^3$ ) is formed? (Assume the reaction goes to completion and all gases are measured at the same temperature and pressure.)

- A. 5.0
- B. 4.0
- C. 3.0
- D. 2.0

(Total 1 mark)

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**CH Cherryholmes**



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