

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, but are made of different molecules and are made up of different numbers of particles. (The particles of the gases are different sizes.)
- At standard atmospheric conditions, all gases have the same average kinetic energy.
- At standard atmospheric conditions, all gases have the same average speed of their particles.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three different phases of matter.



Figure 1: Representation of particles in different phases of matter

Model 1 represents the three phases of matter. The particles in the gas phase are widely spaced and move rapidly. The particles in the liquid phase are closer together and move more slowly. The particles in the solid phase are tightly packed and move very slowly.

Key Concepts

1. Model 1 shows the three phases of matter. The particles in the gas phase are widely spaced and move rapidly. The particles in the liquid phase are closer together and move more slowly. The particles in the solid phase are tightly packed and move very slowly.

Table 1: Representation of the three phases of matter

	Gas	Liquid	Solid
How the particles are arranged:	The particles are widely spaced and move rapidly. They are in constant motion and collide with each other and the walls of the container.	The particles are closer together and move more slowly. They are in constant motion and collide with each other and the walls of the container.	The particles are tightly packed and move very slowly. They are in constant motion and collide with each other and the walls of the container.
How the particles move:	The particles move rapidly and randomly in all directions. They have a high average kinetic energy.	The particles move more slowly and randomly in all directions. They have a lower average kinetic energy than the gas phase.	The particles move very slowly and randomly in all directions. They have the lowest average kinetic energy of the three phases.

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Daniel F McAuley



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