

Wireless Desktop 2000

Key Features



1 Advanced Encryption Standard (AES) 128-Bit Encryption

The Wireless Keyboard 2000 features Advanced Encryption Standard (AES), which can help protect your personal and business information by encrypting what you type.

2 Taskbar Favorites for Windows® 7

Easily access programs in the Windows Taskbar with convenient keyboard shortcut keys.

3 Windows Start Button

Press to launch your Start menu and search your PC or the Web if you have Windows Vista or later installed.

4 Pillow-Texture Palm Rest

The pillow-textured palm rest gives comfortable support while you type.

5 Comfort and control for either hand

Experience full-size comfort with enhanced side grips for control, a shape that fits either hand and a Tilt Wheel for easy side-to-side scrolling.

6 BlueTrack Technology®

Works on virtually any surface in your home, office or anywhere in between.

System Requirements

Computer/Operating System

Windows 7, Windows Vista or Windows XP (excluding Windows XP 64-bit)

Hard Drive

100MB

Other

- USB Port
- Microsoft IntelliPoint v8.0 and Microsoft IntelliType Pro v8.0

Internet access required for certain features. Local and/or long-distance telephone call charges may apply.

Software download required for full functionality of all features.

Internet access required for software installation: www.microsoft.com/hardware/en-us/download

You must accept License Terms for software download.

2 AAA alkaline batteries required (included).

Microsoft Keyboard 1000 Manual

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Microsoft Keyboard 1000 Manual:

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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$1\text{K}_3\text{PO}_4 + \text{Al}(\text{NO}_3)_3 + 3\text{KNO}_3 \rightarrow 3\text{KNO}_3 + \text{AlPO}_4 + 3\text{KNO}_3$ net ionic $\text{Al}^{3+} + \text{PO}_4^{3-} \rightarrow \text{AlPO}_4$ s 2 2 fec13 aq 3 mg s

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^-aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$ because two NH_4^+aq

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