

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

Limiting Reactant And Limiting Yield Answer Key

Rachel Sandford



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Table of Contents Limiting Reactant And Limiting Yield Answer Key

1. Understanding the eBook Limiting Reactant And Limiting Yield Answer Key
 - The Rise of Digital Reading Limiting Reactant And Limiting Yield Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactant And Limiting Yield Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactant And Limiting Yield Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactant And Limiting Yield Answer Key
 - Personalized Recommendations
 - Limiting Reactant And Limiting Yield Answer Key User Reviews and Ratings
 - Limiting Reactant And Limiting Yield Answer Key and Bestseller Lists
5. Accessing Limiting Reactant And Limiting Yield Answer Key Free and Paid eBooks

- Limiting Reactant And Limiting Yield Answer Key Public Domain eBooks
 - Limiting Reactant And Limiting Yield Answer Key eBook Subscription Services
 - Limiting Reactant And Limiting Yield Answer Key Budget-Friendly Options
6. Navigating Limiting Reactant And Limiting Yield Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactant And Limiting Yield Answer Key Compatibility with Devices
 - Limiting Reactant And Limiting Yield Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactant And Limiting Yield Answer Key
 - Highlighting and Note-Taking Limiting Reactant And Limiting Yield Answer Key
 - Interactive Elements Limiting Reactant And Limiting Yield Answer Key
 8. Staying Engaged with Limiting Reactant And Limiting Yield Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactant And Limiting Yield Answer Key
 9. Balancing eBooks and Physical Books Limiting Reactant And Limiting Yield Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactant And Limiting Yield Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limiting Reactant And Limiting Yield Answer Key
 - Setting Reading Goals Limiting Reactant And Limiting Yield Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limiting Reactant And Limiting Yield Answer Key
 - Fact-Checking eBook Content of Limiting Reactant And Limiting Yield Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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