

# SIGNIFICANT FIGURE WORKSHEET-ANSWER KEY

1) Indicate the number of significant figures in the following:

- |               |          |                |          |                          |          |
|---------------|----------|----------------|----------|--------------------------|----------|
| a) 2.508 L    | <u>4</u> | d) 12 in/1 ft  | <u>∞</u> | g) 29 mm                 | <u>2</u> |
| b) 0.00210 km | <u>3</u> | e) 20 students | <u>∞</u> | h) $3.50 \times 10^3$ nm | <u>3</u> |
| c) 3000 g     | <u>1</u> | f) 45.00 g     | <u>4</u> | i) 8.34±5 g/mL           | <u>3</u> |

2) Round each of the following numbers to the indicated number of significant figures:

- |                                   |                            |                                       |  |
|-----------------------------------|----------------------------|---------------------------------------|--|
| a) 96.528 cm <sup>2</sup> to 3 SF | <u>96.5 cm<sup>2</sup></u> | d) 0.036559 m to 4 SF                 | <u>0.03656 m</u>                         |
| b) 1275 mL to 2 SF                | <u>1300 mL</u>             | e) 0.14500001 lb to 2 SF              | <u>0.15 lb</u>                           |
| c) 26.485 g to 4 SF               | <u>26.48 g</u>             | f) $6.0221367 \times 10^{23}$ to 4 SF | <u><math>6.022 \times 10^{23}</math></u> |

3) Perform the following calculations, giving the answer properly rounded to the correct number of significant figures.

- |                            |                |  |               |
|----------------------------|----------------|--|---------------|
| a) 65g + 25.86g + 145.1g = | <u>236 g</u>   | c) $2.5 \times 10^2 + 2.6 \times 10^3 + 3.4 \times 10^{-1} =$      | <u>2900</u>   |
| b) 415.1g - 25.56g =       | <u>389.5 g</u> | d) $\frac{4.582 \text{ g} + 4.54 \text{ g} + 4.60 \text{ g}}{3} =$ | <u>4.57 g</u> |

4) What is the assumed absolute uncertainty (au) of the following?

- |                          |                             |                         |                 |
|--------------------------|-----------------------------|-------------------------|-----------------|
| a) 0.0770 g              | <u>± 0.0001 g</u>           | d) $1.56 \times 10^3$ m | <u>± 10 m</u>   |
| b) 450 mL                | <u>± 10 mL</u>              | e) 4500 km              | <u>± 100 km</u> |
| c) 35±14 cm <sup>2</sup> | <u>± 100 cm<sup>2</sup></u> | f) 25.0 lb              | <u>± 0.1 lb</u> |

5) What is the assumed relative uncertainty (ru) of the following?

- |                          |              |                         |              |
|--------------------------|--------------|-------------------------|--------------|
| a) 0.0770 g              | <u>0.001</u> | d) $1.56 \times 10^3$ m | <u>0.006</u> |
| b) 450 mL                | <u>0.02</u>  | e) 4500 km              | <u>0.02</u>  |
| c) 35±14 cm <sup>2</sup> | <u>0.03</u>  | f) 25.0 lb              | <u>0.004</u> |

6) Perform the following calculations, giving the answer properly rounded to the correct number of significant figures.

- |   |                |  |  |
|---|----------------|--|--|
| a) $6.72 \text{ m} \times \frac{1 \text{ mm}}{10^{-3} \text{ m}} =$ | <u>6720 mm</u> | c) $\frac{(8.015 \times 10^{-3})(0.550)}{1.9 \times 10^3} =$ | <u><math>2.3 \times 10^{-6}</math></u> |
| b) $\frac{60.00 \times 35}{0.0050} =$                               | <u>420,000</u> | d) (22.5 cm)(125.6 cm)(1.4 cm) =                             | <u>40±00 cm<sup>3</sup></u>            |

7) Perform the following calculations, giving the answer properly rounded to the correct number of significant figures. Hint: You might want to show an intermediate answer.

- a)  $\frac{35.46 \text{ g/mol} - 35.4527 \text{ g/mol}}{35.4527 \text{ g/mol}} \times 100\% =$  0.02±1 %

# Mastering Chemistry Answer Key Significant Figures

**Anthony S. Fauci**



## Mastering Chemistry Answer Key Significant Figures:

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