

Limiting and Excess Reactants

Answers

- ① For the following equation, determine which reactant is limiting and which reactant is in excess. The amount of reagents used is shown.



$$40 \text{ g Fe} \times \frac{1 \text{ mol Fe}}{55.8 \text{ g Fe}} \times \frac{1 \text{ mol Fe}_3\text{O}_4}{3 \text{ mol Fe}} = 0.239 \text{ mol Fe}_3\text{O}_4$$

$$16 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Fe}_3\text{O}_4}{4 \text{ mol H}_2\text{O}} = 0.222 \text{ mol Fe}_3\text{O}_4$$

0.222 mol is less than 0.239 mol. Therefore, the limiting reactant is H₂O, and the excess reactant is Fe.

- ② 35.5 g of silver nitrate is reacted with 35.5 grams of sodium sulfide, producing silver sulfide and sodium nitrite.

- Ⓐ Write and balance the equation.



- Ⓑ Calculate the number of grams of silver sulfide produced in grams.

$$35.5 \text{ g AgNO}_3 \times \frac{1 \text{ mol AgNO}_3}{153.9 \text{ g AgNO}_3} \times \frac{1 \text{ mol Ag}_2\text{S}}{2 \text{ mol AgNO}_3} = 0.115 \text{ mol Ag}_2\text{S}$$

$$35.5 \text{ g Na}_2\text{S} \times \frac{1 \text{ mol Na}_2\text{S}}{78 \text{ g Na}_2\text{S}} \times \frac{1 \text{ mol Ag}_2\text{S}}{2 \text{ mol Na}_2\text{S}} = 0.455 \text{ mol Ag}_2\text{S}$$

The amount of silver sulfide produced is 0.115 mol. Let us convert this into grams.

$$0.115 \text{ mol Ag}_2\text{S} \times \frac{247.8 \text{ g Ag}_2\text{S}}{1 \text{ mol Ag}_2\text{S}} = 28.5 \text{ g Ag}_2\text{S}$$

- ③ For the following reaction: $\text{C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$

- Ⓐ Write and balance the equation.

$$0.37 \text{ g CO}_2 \times \frac{1 \text{ mol of CO}_2}{44.01 \text{ g CO}_2} \times \frac{1 \text{ mol C}_3\text{H}_8}{3 \text{ mol CO}_2} = 0.00280 \text{ mol C}_3\text{H}_8$$

- Ⓑ How many grams of propane is this.

$$0.00280 \text{ mol C}_3\text{H}_8 \times \frac{44.1 \text{ g C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} = 0.12348 \text{ g C}_3\text{H}_8$$

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DP Hallahan



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Table of Contents Limited And Excess Reactants Pogil Answers

1. Understanding the eBook Limited And Excess Reactants Pogil Answers
 - The Rise of Digital Reading Limited And Excess Reactants Pogil Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Limited And Excess Reactants Pogil Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limited And Excess Reactants Pogil Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limited And Excess Reactants Pogil Answers

- Personalized Recommendations
 - Limited And Excess Reactants Pogil Answers User Reviews and Ratings
 - Limited And Excess Reactants Pogil Answers and Bestseller Lists
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 - Limited And Excess Reactants Pogil Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Limited And Excess Reactants Pogil Answers Compatibility with Devices
 - Limited And Excess Reactants Pogil Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limited And Excess Reactants Pogil Answers
 - Highlighting and Note-Taking Limited And Excess Reactants Pogil Answers
 - Interactive Elements Limited And Excess Reactants Pogil Answers
 8. Staying Engaged with Limited And Excess Reactants Pogil Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limited And Excess Reactants Pogil Answers
 9. Balancing eBooks and Physical Books Limited And Excess Reactants Pogil Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limited And Excess Reactants Pogil Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limited And Excess Reactants Pogil Answers
 - Setting Reading Goals Limited And Excess Reactants Pogil Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limited And Excess Reactants Pogil Answers

- Fact-Checking eBook Content of Limited And Excess Reactants Pogil Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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