

Limiting Reactant

For the following reactions, find the following:

- Which of the reactants is the limiting reactant?
 - What is the maximum amount of each product that can be formed?
 - How much of the excess reactant is left over after the reaction is complete?
- See next page for calculations!

1. Consider the following reaction:



a) limiting reactant

Answer the questions above, assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate.

$$\begin{aligned} & \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 18.6 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\ & \frac{50.0 \text{ g Na}_3\text{PO}_4}{1} \times \frac{1 \text{ mol Na}_3\text{PO}_4}{162.94 \text{ g Na}_3\text{PO}_4} \times \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol Na}_3\text{PO}_4} \times \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} = 45.5 \text{ g } (\text{NH}_4)_3\text{PO}_4 \\ & \frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \times \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \times \frac{3 \text{ mol NaNO}_3}{3 \text{ mol NH}_4\text{NO}_3} \times \frac{85.00 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 31.9 \text{ g NaNO}_3 \end{aligned}$$

2. Consider the following reaction:



a) limiting reactant

Answer the questions at the top of this sheet, assuming we start with 100.0 grams of calcium carbonate and 45.0 grams of iron (II) phosphate.

$$\begin{aligned} & \frac{100.0 \text{ g CaCO}_3}{1} \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCO}_3} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 103.3 \text{ g Ca}_3(\text{PO}_4)_2 \\ & \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol FePO}_4} \times \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 46.3 \text{ g Ca}_3(\text{PO}_4)_2 \\ & \frac{45.0 \text{ g FePO}_4}{1} \times \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \times \frac{1 \text{ mol Fe}_2(\text{CO}_3)_3}{2 \text{ mol FePO}_4} \times \frac{291.73 \text{ g Fe}_2(\text{CO}_3)_3}{1 \text{ mol Fe}_2(\text{CO}_3)_3} = 43.5 \text{ g Fe}_2(\text{CO}_3)_3 \end{aligned}$$

Limiting Reactant Review Sheet Key

Shasha Hu



Limiting Reactant Review Sheet Key:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises **CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE** Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text **Introduction to General, Organic, and Biochemistry** Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career **Study Guide for Introductory Chemistry** Iris Stovall, 1990 *Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They'll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

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General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999

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Holt Chemistry Ralph Thomas Myers, 2004

Student Study Guide to accompany

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