

Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent

$$14.8 \text{ g C}_3\text{H}_8 \times \frac{1 \text{ mol C}_3\text{H}_8}{44.09 \text{ g C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

O_2 is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.84 \text{ g H}_2\text{O}$$

c) determine the number of grams of H_2O produced

$$1.55 \text{ g H}_2\text{O}$$

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{1 \text{ mol C}_3\text{H}_8}{5 \text{ mol O}_2} \times \frac{44.09 \text{ g C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} = .948 \text{ g C}_3\text{H}_8 \text{ used}$$

$$14.8 \text{ g C}_3\text{H}_8 - 0.948 \text{ g C}_3\text{H}_8 = 13.9 \text{ g C}_3\text{H}_8 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{294.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g NaOH} \times \frac{1 \text{ mol NaOH}}{294.17 \text{ g NaOH}} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 15.8 \text{ g Na}_2\text{SO}_4$$

Limiting Reactions 12 3 Answers

**Jason Chin, Vault Reports (Firm), Job
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Journal of the American Medical Association American Medical Association,1925

Sunday School Times ,1893

Study Guide for the Core Curriculum for Oncology Nursing Suzanne M. Mahon,2005 Published in collaboration with the ONS this Study Guide is a must have for those who are taking the certification exam The chapters parallel those presented in the Core Curriculum for Oncology Nursing 4e The number of questions in each chapter will correspond with the percentage of questions on that particular topic included in the certification exam Answers and rationales for correct and incorrect responses are listed at the end of each chapter The book also includes an updated bibliography for each subject

Teaching Science for Understanding James Joseph Gallagher,2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

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