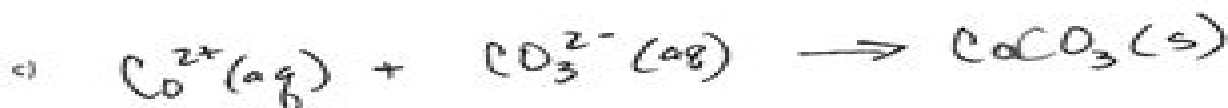
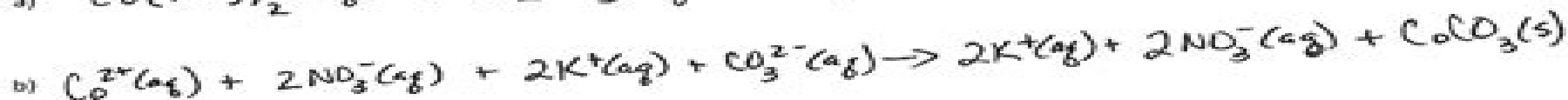
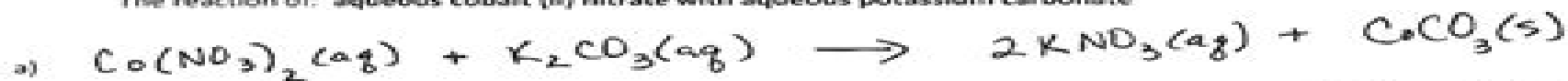


# Names \_\_\_\_\_ Answer Key

1. Write the a) balanced molecular equation, b) the balanced total ionic equation and the c) balanced net ionic equation for the following reaction. Include phases.

The reaction of: aqueous cobalt (II) nitrate with aqueous potassium carbonate



d) Circle all the descriptors below that describe the reaction:

precipitation

gas-forming

redox

double-replacement

acid-base

single-replacement

e) List all acids in the reaction (if any):

NONE

f) List all bases in the reaction (if any):

$\text{K}_2\text{CO}_3$ ,  $\text{CoCO}_3$

g) List all precipitates (if any):

$\text{CoCO}_3$

h) What is the oxidation number of the carbon in potassium carbonate?

$\text{CO}_3^{2-}$

$\text{C} + 3(-2) = -2$

$\text{C} = +4$

i) What element (if any) was oxidized?

NONE

# Lab 2bags Of Reactions Answers

**McGraw-Hill Staff**



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