

## ELECTRON CONFIGURATIONS OF ATOMS AND IONS

### OBJECTIVES

- Observe colors of various chemical solutions
- Write electron configurations for various metals, nonmetals, cations, and anions
- Relate the presence of color in an ionic solution as a characteristic of electron configurations

### PURPOSE

In this experiment, you will observe a variety of chemical solutions containing common cations and anions. You will write electron configurations for many of the ions contained in the solutions. You will observe the colored solutions and draw conclusions about the electron configurations of the metal ions in the colored solutions.

### EXPERIMENTAL PAGE

1. Place one drop of each of the indicated solutions on the acetate sheet. Record the color of each solution

NaCl	MgSO <sub>4</sub>	AlCl <sub>3</sub>
FeCl <sub>3</sub>	CaCl <sub>2</sub>	NiSO <sub>4</sub>
CuSO <sub>4</sub>	ZnCl <sub>2</sub>	AgNO <sub>3</sub>

2. A precipitate is a solid that separates upon mixing solutions. Predict which of the metal cations in this experiment will form colored precipitates upon the addition of NaOH. **That prediction is based on reading the introduction.** Add one drop of NaOH to find out. Record your results.

NaCl	MgSO <sub>4</sub>	AlCl <sub>3</sub>
FeCl <sub>3</sub>	CaCl <sub>2</sub>	NiSO <sub>4</sub>
CuSO <sub>4</sub>	ZnCl <sub>2</sub>	AgNO <sub>3</sub>

# Lab 10 Electron Configurations Of Atoms And Ions

**Electrochemical Society**



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*Laboratory Experiments for Foundations of Chemistry* Ernest R. Toon, 1973      *Laboratory Manual for Chemistry* Karen Timberlake, 1979      **Contributions from the Physical Laboratory of the University of Michigan** University of Michigan Physical Laboratory, 1913      *Nuclear Science Abstracts* , 1968      *Technical Paper - Florida Engineering and Industrial Experiment Station* University of Florida. Engineering and Industrial Experiment Station, 1960      **Laboratory Exercises for Preparatory Chemistry** Kathy Dodds Tyner, 1994-06 Laboratory Exercises for Preparatory Chemistry is the perfect complement to a one semester preparatory chemistry laboratory course Tyner s manual emphasizes the application of chemistry and the principles of science to everyday life The labs are directly applicable to the real world and often contain supplemental assignments that illustrate an application      **Hygienic Laboratory bulletin. no. 115-120, 1918-20** , 1919      *Using Physical Models of Biomolecules to Teach Concepts of Biochemical Structure in Introductory Undergraduate Chemistry* John Yi, 2004      **Stanford Synchrotron Radiation Laboratory Activity Report 7/1/77 - 3/31/78** Stanford Synchrotron Radiation Laboratory. Stanford Synchrotron Radiation Laboratory, Arthur Bienenstock, 1978      **International Bulletin of Information on Refrigeration** International Institute of Refrigeration, 1929      **Publications** General Electric Company. Research Laboratory, 1919      Transactions Electrochemical Society, 1929      *Transactions of the Electrochemical Society* Electrochemical Society, 1929      **Transactions of the American Electrochemical Society** , 1929      **Chemistry and Industry Review** , 1928      Calendar University of Saskatchewan, 1925      **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1928      Advanced Laboratory Practice in Electricity and Magnetism Earle Melvin Terry, 1929 Batcheller Collection      *Contributions from the Physical Laboratory* University of Michigan. Department of Physics, 1925      **Researches of the Electrotechnical Laboratory** Tokyo. Electrotechnical Laboratory, 1928

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