

Empirical and Molecular Formulas Worksheet

Key

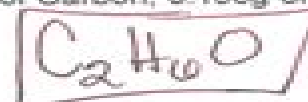
Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} = 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} = 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1 \end{array}$$



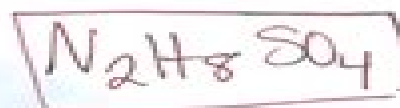
- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{array}{l} \text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \\ \text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5 \end{array}$$



- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\begin{array}{l} \text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2 \\ \text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8 \\ \text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1 \\ \text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4 \end{array}$$



Molecular Formula

- 4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) = 30$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

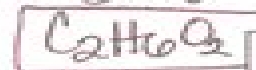
$$\text{C} \quad \frac{1.388}{3.585} \rightarrow 38.7\% \rightarrow \frac{3.87\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \rightarrow 9.62\% \rightarrow \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \rightarrow 51.6\% \rightarrow \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	NH_4SO_4
4	$\text{C}_2\text{H}_6\text{O}_2$
5	$\text{C}_6\text{H}_{12}\text{O}_6$

Lab 6 Empirical Formula Answers

LL Leslie

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the LL Leslie text.

Lab 6 Empirical Formula Answers:

Contemporary Chemistry Leonard Saland, 1993 This comprehensive guide gives you lesson plans activities and tests for two sequential semester long chemistry courses It is designed to work with our student book Contemporary Chemistry Each lesson plan features a DO NOW section to engage students as soon as they get to class instructional objectives an aim for that class period a motivational application questions or demonstrations to help students draw valid conclusions homework assignments You also get term calendars weekly tests and complete answer keys **S.Chand S Mathematics For Class X**

Term -I H.K. Dass, Rama Verma & Bhagwat S. Sharma, S Chand s Mathematics books for Classes IX and X are completely based on CCE pattern of CBSE The book for Term I covers the syllabus from April to September and the book for Term II covers the syllabus from October to March Pharmaceutical Calculations Michalakakis Savva, 2019-10-17 Pharmaceutical Calculations A Conceptual Approach is a book that combines conceptual and procedural understanding for students and will guide you to master prerequisite skills to carry out accurate compounding and dosage regimen calculations It is a book that makes the connection between basic sciences and pharmacy It describes the most important concepts in pharmaceutical sciences thoroughly accurately and consistently through various commentaries and activities to make you a scientific thinker and to help you succeed in college and licensure exams Calculation of the error associated with a dose measurement can only be carried out after understanding the concept of accuracy versus precision in a measurement Similarly full appreciation of drug absorption and distribution to tissues can only come about after understanding the process of transmembrane passive diffusion Early understanding of these concepts will allow reinforcement and deeper comprehension of other related concepts taught in other courses More weight is placed on the qualitative understanding of fundamental concepts like tonicity vs osmotic pressure diffusion vs osmosis crystalloids vs colloids osmotic diuretics vs plasma expanders rate of change vs rate constants drug accumulation vs drug fluctuation loading dose vs maintenance dose body surface area BSA vs body weight BW as methods to adjust dosages and much more before considering other quantitative problems In one more significant innovation the origin and physical significance of all final forms of critical equations is always described in detail thus allowing recognition of the real application and limitations of an equation Specific strategies are explained step by step in more than 100 practice examples taken from the fields of compounding pharmacy pharmaceuticals pharmacokinetics pharmacology and medicine **Million Dollar Data: Building Confidence - Vol.1** Stephen DeMeo, 2020-07-31 Global

warming our current and greatest challenge is without precedent Among the many consequences that are impacting our society one unanticipated concern involves scientific truth When the President of the United States and others in his administration declare that global warming is fake science it calls into question what real science is and what real school science should be I will argue that real science is quality science one that is based on the rigorous collection of reliable and valid data To collect quality data requires bending over backwards to get things right and this is exactly what makes science

so special Truth is made when scientists go this extra yard and devise controlled experiments collect large data sets confirm the data and rationally analyze their results Making scientific truth sounds difficult to do in the science laboratory but in reality there are many straightforward ways that truth can be constructed In the first of two volumes I discuss twelve such ways I call them Confidence Indicators that can allow students to strongly believe in their data and their subsequent results Many of these methods are intuitive and can be used by young students on the late elementary level all the way up to those taking introductory college science courses As in life science is not without doubt In the second volume I introduce the concept of scientific uncertainty and the indicators used to calculate its magnitude I will show that science is about connecting confidence with uncertainty in a specific manner what I refer to as the Confidence Uncertainty Continuum expression This important relationship epitomizes the scientific enterprise as a search for probabilistic rather than absolute truth This two volume set will contain a variety of ways that data quality can be instituted into a science curriculum To support its use many of the examples that I will present involve science teachers as well as student work and feedback from different grade levels and in different scientific disciplines Specific chapters will be devoted to reviewing the academic literature on data quality as well as describing my own personal research on this important but often neglected topic E3 Chemistry Review Book - 2018 Home Edition (Answer Key Included) Effiong Eyo, 2017-10-20 With Answer Key to All Questions Chemistry students and homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Review Book 2018 With E3 Chemistry Review Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with solutions to study and follow Several practice multiple choice and short answer questions at the end of each lesson to test understanding of the materials 12 topics of Regents question sets and 3 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN 978 197836229 The Home Edition contains an answer key section Teachers who want to recommend our Review Book to their students should recommend the Home Edition Students and parents whose school is not using the Review Book as instructional material as well as homeschoolers should buy the Home Edition The School Edition does not have answer key in the book A separate answer key booklet is provided to teachers with a class order of the book Whether you are using the school or Home Edition our E3 Chemistry Review Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year PLEASE NOTE Although reading contents in both the school and home editions are identical there are slight differences in question numbers choices and pages between the two editions Students whose school is using the Review Book as instructional material SHOULD NOT buy the Home Edition Also available in paperback print Stoichiometry Unit Project Luann Marie

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Table of Contents Lab 6 Empirical Formula Answers

1. Understanding the eBook Lab 6 Empirical Formula Answers
 - The Rise of Digital Reading Lab 6 Empirical Formula Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab 6 Empirical Formula Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 6 Empirical Formula Answers
 - Personalized Recommendations
 - Lab 6 Empirical Formula Answers User Reviews and Ratings
 - Lab 6 Empirical Formula Answers and Bestseller Lists

5. Accessing Lab 6 Empirical Formula Answers Free and Paid eBooks
 - Lab 6 Empirical Formula Answers Public Domain eBooks
 - Lab 6 Empirical Formula Answers eBook Subscription Services
 - Lab 6 Empirical Formula Answers Budget-Friendly Options
6. Navigating Lab 6 Empirical Formula Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab 6 Empirical Formula Answers Compatibility with Devices
 - Lab 6 Empirical Formula Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 6 Empirical Formula Answers
 - Highlighting and Note-Taking Lab 6 Empirical Formula Answers
 - Interactive Elements Lab 6 Empirical Formula Answers
8. Staying Engaged with Lab 6 Empirical Formula Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 6 Empirical Formula Answers
9. Balancing eBooks and Physical Books Lab 6 Empirical Formula Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 6 Empirical Formula Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab 6 Empirical Formula Answers
 - Setting Reading Goals Lab 6 Empirical Formula Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab 6 Empirical Formula Answers
 - Fact-Checking eBook Content of Lab 6 Empirical Formula Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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