

## Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of  $\text{C}_3\text{H}_8$  and 3.44 g of  $\text{O}_2$ , determine the limiting reactant.
- Determine the number of moles of carbon dioxide produced.
- Determine the number of grams of  $\text{H}_2\text{O}$  produced.
- Determine the number of grams of excess reactant left.

2. Given the following equation:



- If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of  $\text{NaOH}$ , determine the limiting reactant.
- Determine the number of moles of  $\text{Al}(\text{OH})_3$  produced.
- Determine the number of grams of  $\text{Na}_2\text{SO}_4$  produced.
- Determine the number of grams of excess reactant left over in the reaction.

3. Given the following equation:



- If 25.4 g of  $\text{Al}_2\text{O}_3$  is reacted with 10.2 g of  $\text{Fe}$ , determine the limiting reactant.
- Determine the number of moles of  $\text{Al}$  produced.
- Determine the number of grams of  $\text{Fe}_3\text{O}_4$  produced.
- Determine the number of grams of excess reactant left over in the reaction.

4. Consider the reaction



- 80.0 grams of  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ . Determine the mass of iodine  $\text{I}_2$ , which could be produced?
- If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.
  - What mass of iodine was produced?
  - What percentage yield of iodine was produced?

5. Zinc and sulfur react to form zinc sulfide according to the equation



If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- Which chemical is the limiting reactant?
- How many grams of  $\text{ZnS}$  will be formed?
- How many grams of the excess reactant will remain after the reaction is over?

6. Which element is in excess when 3.00 grams of  $\text{Mg}$  is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of  $\text{MgO}$  is formed?

7. How many grams of  $\text{Al}_2\text{S}_3$  are formed when 5.00 grams of  $\text{Al}$  is heated with 10.0 grams  $\text{S}$ ?

# Limiting Reagent Practice 2 Answers

**Kenneth W. Raymond**



## Limiting Reagent Practice 2 Answers:

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