

Final Exam Review Multiple-Choice Questions Combined

- Use the VSEPR theory to predict molecular geometry of BrF_3 .
 a) tetrahedral
 b) see-saw
 c) trigonal bipyramidal
 d) square pyramidal
 e) octahedral
- Use the VSEPR theory to predict molecular geometry of CO_3^{2-} .
 a) bent
 b) tetrahedral
 c) trigonal pyramidal
 d) T-shaped
 e) trigonal planar
- What is the bond angle between the O-N-O bonds in NO_2^+ ion?
 a) 180°
 b) 120°
 c) 118°
 d) 90°
 e) 80°
- What is the molecular shape of SeO_4^{2-} ?
 a) see-saw
 b) trigonal pyramidal
 c) trigonal planar
 d) tetrahedral
 e) T-shaped
- What is the molecular shape of the SF_6^- ion?
 a) linear
 b) bent
 c) trigonal planar
 d) tetrahedral
 e) T shaped
- For XeF_4 , the valence electron geometry around the central atom is _____.
 a) trigonal planar
 b) tetrahedral
 c) trigonal bipyramidal
 d) octahedral
 e) square planar
- The molecular shape for XeF_4 is _____.
 a) trigonal planar
 b) tetrahedral
 c) trigonal bipyramidal
 d) octahedral
 e) square planar
- The F-Xe-F bond angle in XeF_4 molecule is _____.
 a) 90°
 b) 109.5°
 c) none of these
 d) 180°

For the following 4 questions, the following molecular substances are referred to. Any substance may be used more than once in answering the questions. CO_2 , SO_2 , SiH_4 , XeF_4 .

9. The species which has tetrahedral as the valence electron geometry is _____.
a) CO_2 c) SiF_4 e) none of these
b) SO_2 d) XeF_4 f) these
10. The species which has linear molecular geometry is _____.
a) CO_2 c) SiF_4 e) none of these
b) SO_2 d) XeF_4 f) these
11. The species which is polar is _____.
a) CO_2 c) SiF_4 e) none of these
b) SO_2 d) XeF_4 f) these
12. The species which has two lone pairs of valence electrons on the central atom is _____.
a) CO_2 b) SO_2 c) SiF_4 d) XeF_4

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J Rink



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