

# MASTER FORMULA SHEET

## CHAPTER 10:

PV = nRT (IDEAL GAS)

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \text{ (COMBINED GAS LAW)}$$

$$J = \frac{C_M}{RT} \text{ (CALCULATING DEVIATION OR MOLAL RATIO)}$$

$$P_1 = P_2 = P_3 = \dots \text{ (TOTAL PRESSURE OR PARTIAL PRESSURES)}$$

$$P_1 = \left(\frac{n_1}{n_2}\right) P_2 = X_1 P_2 \text{ (RELATING T. PRESS. TO MOLE FRACTION)}$$

$$M_{rms} = \sqrt{\frac{3RT}{M}} \text{ (DEFINITION OF R.M.S. VELOCITY)}$$

$$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}} \text{ (RATE OF EFFUSION TO MOLE WEIGHTS)}$$

## CHAPTER 11: IN INTERMOLECULAR FORCES

LPE < DIPOLE-DIPOLE < H-BONDING < VAN DER WAALS

IF WEIGHT OF A MOLECULE IS MUCH > ANOTHER, DISPERSIONAL FORCES

$$\ln P = \frac{-\Delta H_{vap}}{RT} + C \text{ (CLAUDETT-CLAUDEMAN)}$$

WITH GRAPH:  $\Delta H_{vap}$  C-SLOPE = R

## CHAPTER 12:

REC. CAT. / FORM. UNIT = AN. COORD. N

REC. CAT. / FORM. UNIT = AN. COORD. N

PRIM. CUBE

REC. CENT. LATT.

FAC. CENT. LATTICE

FAC. CENT. LATTICE

(LATTICE / 2 ATOMS) (4 ATOMS) (4 ATOMS) (VOL. OF SPHERE)

## CHAPTER 15:

$$K_c = \frac{[D]^d [E]^e}{[A]^a [B]^b}$$

(USES ONLY TO CONC.)

$$K_p = \frac{(P_D)^d (P_E)^e}{(P_A)^a (P_B)^b}$$

$$K_p = K_c (RT)^{\Delta n}$$

$$Q_c = \frac{[D]^d [E]^e}{[A]^a [B]^b}$$

THE REQ. QUOTIENT. @ ANY TIME IN RXN IF @ EQ.  $Q_c = K_c$

EQUILIBRIUM CONSTANT

RELATING THE TWO EQ'S

RELATING THE TWO EQ'S

RELATING THE TWO EQ'S

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## CHAPTER 17:

$$\Delta S = 4.184/T \text{ (}\Delta S \text{ IN J/K)}$$

2ND LAW OF THERMODYNAMICS

$$\Delta S_{univ} = \Delta S_{sys} + \Delta S_{surr} = 0$$

$$\Delta S_{univ} = \Delta S_{sys} + \Delta S_{surr} > 0$$

$$\Delta S = \int \frac{dq_{rev}}{T}$$

$$\Delta S_{surr} = \frac{-\Delta H_{sys}}{T}$$

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## CHAPTER 16:

$$K_w = [H_3O^+][OH^-] = 1.0 \times 10^{-14}$$

$$pH = -\log[H^+] \text{ (THIS IS @ 25°C)}$$

$$pOH = -\log[OH^-]$$

$$pH + pOH = 14.00$$

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} \text{ OR } \frac{[H^+][A^-]}{[HA]}$$

$$\text{PERCENT IONIZATION} = \frac{[H_3O^+][A^-]}{[HA]_{\text{INIT.}}} \times 100$$

$$K_b = \frac{[OH^-][OH^-]}{[B]} \quad K_a \cdot K_b = K_w$$

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$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$\Delta G^\circ = \sum \Delta G_f^\circ (\text{PROD}) - \sum \Delta G_f^\circ (\text{REACT})$$

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# Maddahmik Writeing Formola

**Liyng Dong**



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