

1. As a solute is added to a solvent, what happens to the freezing point and the boiling point of the solution?
- ☒ A) The freezing point decreases and the boiling point increases.
- B) The freezing point increases and the boiling point increases.
- C) The freezing point increases and the boiling point decreases.
- D) The freezing point decreases and the boiling point decreases.
2. Compared to the freezing point of 1.0 M KCl(aq) at standard pressure, the freezing point of 1.0 M CaCl<sub>2</sub>(aq) at standard pressure is
- More ions*  
☒ A) lower  
 B) higher  
 C) the same
3. Which aqueous solution of KI freezes at the lowest temperature?
- A) 1 mol of KI in 500. g of water
- ☒ B) 2 mol of KI in 500. g of water *(more concentrated)*
- C) 1 mol of KI in 1000. g of water
- D) 2 mol of KI in 1000. g of water
4. A 1 kilogram sample of water will have the highest freezing point when it contains
- A)  $1 \times 10^{21}$  dissolved particles
- ☒ B)  $1 \times 10^{17}$  dissolved particles
- C)  $1 \times 10^{19}$  dissolved particles
- D)  $1 \times 10^{23}$  dissolved particles *lowest f.p.*
5. A solution consists of 0.50 mole of CaCl<sub>2</sub> dissolved in 100. grams of H<sub>2</sub>O at 25°C. Compared to the boiling point and freezing point of 100. grams of H<sub>2</sub>O at standard pressure, the solution at standard pressure has
- ☒ A) a higher boiling point and a lower freezing point
- B) a lower boiling point and a lower freezing point
- C) a higher boiling point and a higher freezing point
- D) a lower boiling point and a higher freezing point
6. Which solution will freeze at the lowest temperature?
- A) 1 mole of sugar in 500 g of water
- B) 1 mole of sugar in 1,000 g of water
- C) 2 moles of sugar in 1,000 g of water
- ☒ D) 2 moles of sugar in 500 g of water
7. Which solution has the highest boiling point at standard pressure?
- 4 ions 3 K<sup>+</sup> 3 PO<sub>4</sub><sup>-</sup>*  
☒ A) 0.10 M K<sub>3</sub>PO<sub>4</sub>(aq)
- B) 0.10 M KCl(aq)
- C) 0.10 M KNO<sub>3</sub>(aq)
- D) 0.10 M K<sub>2</sub>SO<sub>4</sub>(aq)
8. Which concentration of a solution of CH<sub>3</sub>OH in water has the lowest freezing point?
- A) 0.001 M
- ☒ B) 0.1 M
- C) 0.01 M
- D) 0.0001 M
9. How do the boiling point and freezing point of a solution of water and calcium chloride at standard pressure compare to the boiling point and freezing point of water at standard pressure?
- A) Both the freezing point and boiling point of the solution are higher.
- B) The freezing point of the solution is higher and the boiling point of the solution is lower.
- ☒ C) The freezing point of the solution is lower and the boiling point of the solution is higher.
- D) Both the freezing point and boiling point of the solution are lower.
10. Compared to a 2.0 M aqueous solution of NaCl at 1 atmosphere, a 3.0 M aqueous solution of NaCl at 1 atmosphere has a
- A) lower boiling point and a higher freezing point
- B) higher boiling point and a higher freezing point
- C) lower boiling point and a lower freezing point
- ☒ D) higher boiling point and a lower freezing point

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Earth Science: the Physical Setting Paola Santagostino,Prentice Hall (School Division),2005 Focusing on the Earth Science content tested on the Regents Examination this thorough review guide contains extensive vocabulary review questions and Memory Jogger and Digging Deeper features Hundreds of practice questions organized in the Regents Examination format help students familiarize themselves with look and feel of the actual exam **Cracking the Regents Earth Science, 2000 Edition** Princeton Review Publishing Staff,Kim Magloire,2000-02-15 5 Actual Exams with Answers Explained Plus the August 1999 Exam It s no secret The best way to ace the Regents exam is by practicing on real tests This guide includes 5 actual full length Earth Science Regents exams with answers and complete explanations plus the August 1999 exam In Cracking the Regents Earth Science 2000 Edition the Regents experts at The Princeton Review teach you the test taking techniques you ll need to know Focus on the material that is most likely to show up on the test Use process of elimination to guess when you re not sure of an answer Practice your skills on the actual Earth Science Regents exams inside Visit [www review com regents](http://www.review.com/regents) for the latest Regents updates and for the January 2000 exam Catalog of Copyright Entries ,1918 **Reclaiming Accountability in Teacher Education** Marilyn Cochran-Smith,Molly Cummings Carney,Elizabeth Stringer Keefe,Stephani Burton,Wen-Chia Chang,M. Beatriz Fernández,Andrew F. Miller,Juan Gabriel Sánchez,Megina Baker,2018-04-20 1 The book offers teacher educators and stakeholders an overview of accountability in the era of education reform and embraces teacher education accountability as a lever for reconstructing its targets purposes and consequences in keeping with the larger democratic project 2 The book introduces a framework eight dimensions of accountability for interrogating dimensions of accountability policy and practice by revealing an accountability initiative s operation but also exposing underlying values and principles theory of change and relationship to larger political and policy agendas 3 Using the authors framework eight dimensions of accountability the book deconstructs four of the most visible education reform

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