

- 1 Amended from AQA Chemistry Unit 1 Foundation Chemistry CHEM1 June 2010  
(Question 1)

Fluorine forms many compounds that contain covalent bonds.

- a i State the meaning of the term *covalent bond*.

.....  
.....

(1 mark)

- ii Write an equation to show the formation of one molecule of  $\text{ClF}_3$  from chlorine and fluorine molecules.

.....  
.....

(1 mark)

- b Draw the shape of a dichlorodifluoromethane molecule ( $\text{CCl}_2\text{F}_2$ ) and the shape of a chlorine trifluoride molecule ( $\text{ClF}_3$ ). Include any lone pairs of electrons that influence the shape.

Shape of  $\text{CCl}_2\text{F}_2$

Shape of  $\text{ClF}_3$

(2 marks)

- c Suggest the strongest type of intermolecular force between  $\text{CCl}_2\text{F}_2$  molecules.

.....  
.....

(1 mark)

- d  $\text{BF}_3$  is a covalent molecule that reacts with an  $\text{F}^-$  ion to form a  $\text{BF}_4^-$  ion.

- i Name the type of bond formed when a molecule of  $\text{BF}_3$  reacts with an  $\text{F}^-$  ion. Explain how this bond is formed.

Type of bond

.....  
.....

Explanation

.....  
.....

(2 marks)

# June 2013 Aqa Chem1 Paper

**University of Sydney. Department of  
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### **Table of Contents June 2013 Aqa Chem1 Paper**

1. Understanding the eBook June 2013 Aqa Chem1 Paper
  - The Rise of Digital Reading June 2013 Aqa Chem1 Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Aqa Chem1 Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an June 2013 Aqa Chem1 Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Aqa Chem1 Paper
  - Personalized Recommendations
  - June 2013 Aqa Chem1 Paper User Reviews and Ratings

- June 2013 Aqa Chem1 Paper and Bestseller Lists
- 5. Accessing June 2013 Aqa Chem1 Paper Free and Paid eBooks
  - June 2013 Aqa Chem1 Paper Public Domain eBooks
  - June 2013 Aqa Chem1 Paper eBook Subscription Services
  - June 2013 Aqa Chem1 Paper Budget-Friendly Options
- 6. Navigating June 2013 Aqa Chem1 Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - June 2013 Aqa Chem1 Paper Compatibility with Devices
  - June 2013 Aqa Chem1 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of June 2013 Aqa Chem1 Paper
  - Highlighting and Note-Taking June 2013 Aqa Chem1 Paper
  - Interactive Elements June 2013 Aqa Chem1 Paper
- 8. Staying Engaged with June 2013 Aqa Chem1 Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers June 2013 Aqa Chem1 Paper
- 9. Balancing eBooks and Physical Books June 2013 Aqa Chem1 Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection June 2013 Aqa Chem1 Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 Aqa Chem1 Paper
  - Setting Reading Goals June 2013 Aqa Chem1 Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 Aqa Chem1 Paper
  - Fact-Checking eBook Content of June 2013 Aqa Chem1 Paper
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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